



## **Appendix E.2**

Beaver Dam Project - ML/ARD Assessment Report – Dec. 20, 2018 as  
Completed for the Updated 2021 Beaver Dam Mine EIS



ATLANTIC GOLD

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## ***Beaver Dam Project – ML/ARD Assessment Report***

**Prepared for:**  
**Atlantic Gold Corporation**  
**Suite 3083, 595 Burrard St.**  
**Vancouver, BC, V7X 1L3**

**Prepared by:**  
**Lorax Environmental Services Ltd.**  
**2289 Burrard St.**  
**Vancouver, BC, V6J 3H9**

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## **1. Introduction**

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ATLANTIC GOLD

# **1    *Introduction***

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The Beaver Dam project is a proposed gold mine owned by Atlantic Mining Nova Scotia Corporation (AMNS) who is preparing a revised Environmental Impact Statement (EIS) that will be submitted to the Canadian Environmental Assessment Agency (CEAA) and Nova Scotia Environment (NSE) as part of the project's regulatory requirements. Lorax Environmental Services Ltd. (Lorax) was retained by AMNS to conduct a geochemical study to characterize mine material such as waste rock, ore, tailings and overburden. This study considers material properties that may change water quality attributed to metal leaching and acid rock drainage (ML/ARD). The geochemical characterization of Beaver Dam samples will also support the development of geochemical source terms for water quality modelling.

ML/ARD is typically associated with the weathering of sulphide-bearing geologic materials. While this is a natural process, the exposure of fresh particle surfaces produced by mining activity enhances the reaction rates associated with ML/ARD.

Following the introduction, a brief background of the geology of the area is provided in Section 2. Section 3 describes the methodology related to the sample selection/collection as well as geochemical analyses and Section 4 discusses the analytical results. Conclusions are provided in Section 5.

## **2. Geology**

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## **2 Geology**

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### **2.1 Regional Geology**

Nova Scotia is divided into two terranes along the east-west trending Cobequid-Chedabucto Fault Zone (FSSI, 2015). The Beaver Dam site is within the southern Meguma Terrane, while the Avalon Terrane is located to the north of the fault zone. The Meguma Terrane includes the Cambrian to Ordovician Meguma Group and Late Ordovician to Early Devonian volcanic and sedimentary rocks (FSSI, 2015). After the collision of the Meguma and Avalon terranes in the mid-Devonian period, sedimentary material was deposited over both terranes during the Carboniferous to Early Cretaceous period. The sedimentary units include siliciclastic rocks, calcareous rocks, evaporites, coal, kaolinitic clay and silica sand (FSSI, 2015).

The majority of the gold mineralization occurs within the units of the Meguma Group. The Meguma Group is divided into the Goldenville Formation and the overlying Halifax Formation. The metamorphic facies in both Meguma Group units vary from greenschist to amphibolite facies. The Goldenville Formation is a greywacke unit that is > 5,600m thick, while the Halifax Formation is primarily argillite with an average thickness of approximately 4,400 m. Both the Goldenville Formation and the Halifax Formation are made up of deep marine turbidite deposits.

### **2.2 Site Geology**

The Meguma Group is the dominant unit occurring in the area near the Beaver Dam site. The Goldenville Formation is further subdivided, from oldest to youngest, into the Moose River Member, the Tangier Member and the Taylors Head Member. Claystone and siltstone are present in the Moose River and Tangier Members but are minor in the Taylors Head Member. Overall there is a decrease in the proportion of fine-grained sediments from the oldest to the youngest units within the Goldenville Formation.

The site of the Beaver Dam project is within the northeast trending Moose River-Beaver Dam Anticline. The metamorphic facies in this region are amphibolite to staurolite facies. The main unit at the project site is the Moose River Member of the Goldenville Formation which is comprised of alternating argillite and greywacke units (FSSI, 2015). While many unique lithologies and geologic structures are tracked during core logging, argillite and greywacke along with two interbedded intermediates (argillite- and greywacke-dominated) represent the four major rock types by volume identified on site. In an effort to be consistent with the geological observations from site while simultaneously allowing for a representative and simplified geo-environmental model, these four major rock types are being carried forward in this ML/ARD assessment. A more detailed description of these units is given in Table 3-1.

### ***3. Samples and Analytical Methods***

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### **3 Samples and Analytical Methods**

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#### **3.1 Sample Selection and Collection**

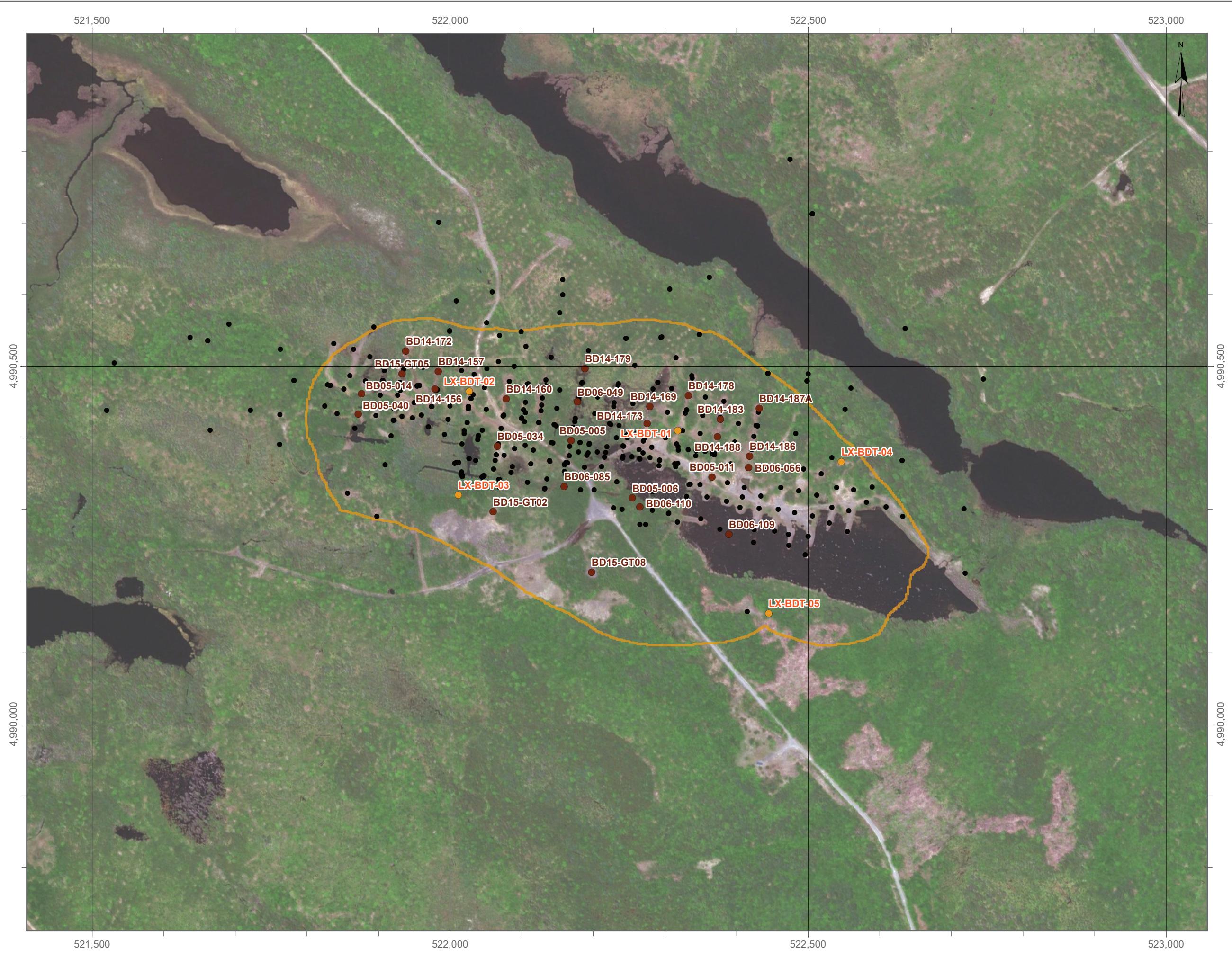
##### **3.1.1 Mine Rock Samples**

A total of 30 Beaver Dam drill core samples were collected by Lorax in November 2017 to supplement the existing mine rock geochemical data set based on 40 drill core samples previously collected by Atlantic Gold in support of the 2017 EIS submission. Static test analyses were carried out on all samples; however, the initial 40 samples were only evaluated for acid base accounting (ABA) characteristics. All Lorax samples were analyzed for ABA and solid phase elements, and based on these results, a subset of samples was selected for kinetic testwork. The kinetic test subsamples were additionally assessed for mineralogy and particle size distribution. The Beaver Dam geologic units and sample descriptions are provided in Table 3-1, whereas the sample locations and drill core details are presented in Appendix 3-1. Eight of the samples collected for the EIS submission could not be classified by lithology as the corresponding logs were not made available. The results from these samples are not included in the report. The location of the sampled drill holes is provided in Figure 3-1.

Eight humidity cells (HC1 through HC8) were initiated using crushed drill core material covering median and high sulphur content ranges for each of the four lithologies, as summarized in Table 3-2. The objective of the program is to provide sulphide oxidation and metal leaching rates for a range of sulphide contents under varying pH conditions to be used as input for the geochemical source term model.

**Table 3-1:**  
**Beaver Dam Geologic Units and Sample Descriptions**

Geologic Unit	Description	Code	No. Samples
Argillite	argillite with < 5% greywacke interbeds	AR	10
Argillite-Greywacke	argillite with 5-49% greywacke interbeds	AG	11
Greywacke-Argillite	greywacke with 20-50% argillite interbeds	GA	14
Greywacke	greywacke with < 20% argillite interbeds	GW	24
Quartz Vein	quartz vein	QTZV	3
Unknown	Core logs were not available	-	8



**Table 3-2:**  
**Summary of humidity cell testwork**

HC ID	Lithology	Rational	Hole ID	Interval (m)		Sample ID
				From	To	
HC1	Argillite	Median sulphur	BD14-178	7	7.9	LX17-12
HC2	Argillite	High sulphur	BD14-188	22	23	LX17-24
			BD15-GT05	57	58	LX17-22
HC3	Argillite-Greywacke	Median sulphur	BD14-172	140	141	LX17-10
			BD15-GT02	15	16	LX17-02
HC4	Argillite-Greywacke	High sulphur	BD15-GT02	26	27	LX17-03
			BD15-GT02	31	32	LX17-04
HC5	Greywacke-Argillite	Median sulphur	BD14-188	10	11	LX17-23
			BD15-GT02	10	11	LX17-01
HC6	Greywacke-Argillite	High sulphur	BD14-178	161	162	LX17-17
			BD15-GT05	15	16	LX17-19
HC7	Greywacke	Median sulphur	BD14-178	49	50	LX17-16
HC8	Greywacke	High sulphur	BD15-GT05	25	26	LX17-20
			BD15-GT05	41	42	LX17-21

### 3.1.2 Overburden Samples

Five overburden samples were collected in September 2018 from within the Beaver Dam pit footprint (Figure 3-1). The samples were collected from shallow (30-45 cm) test pits dug by hand using a shovel. In general, the overburden was silty sand with minor gravel and varying amounts of cobbles (Table 3-3). Cobbles greater than approximately 5 cm in diameter were removed from the sample submitted to the lab. The analyses for these samples included ABA, solid phase elements and shake flask extractions (SFE).

**Table 3-3:**  
**Summary of overburden samples**

Sample ID	Test Pit Depth (cm)	Sample Description
LX-BDT-01	30	Silty sand with some clay, minor gravel and cobbles
LX-BDT-02	36	Silty sand, minor gravel
LX-BDT-03	43	High organic content, silty with some cobbles
LX-BDT-04	33	Silty sand with gravel and several cobbles
LX-BDT-05	34	Silty sand with gravel and cobbles

### 3.1.3 Haul Road Samples

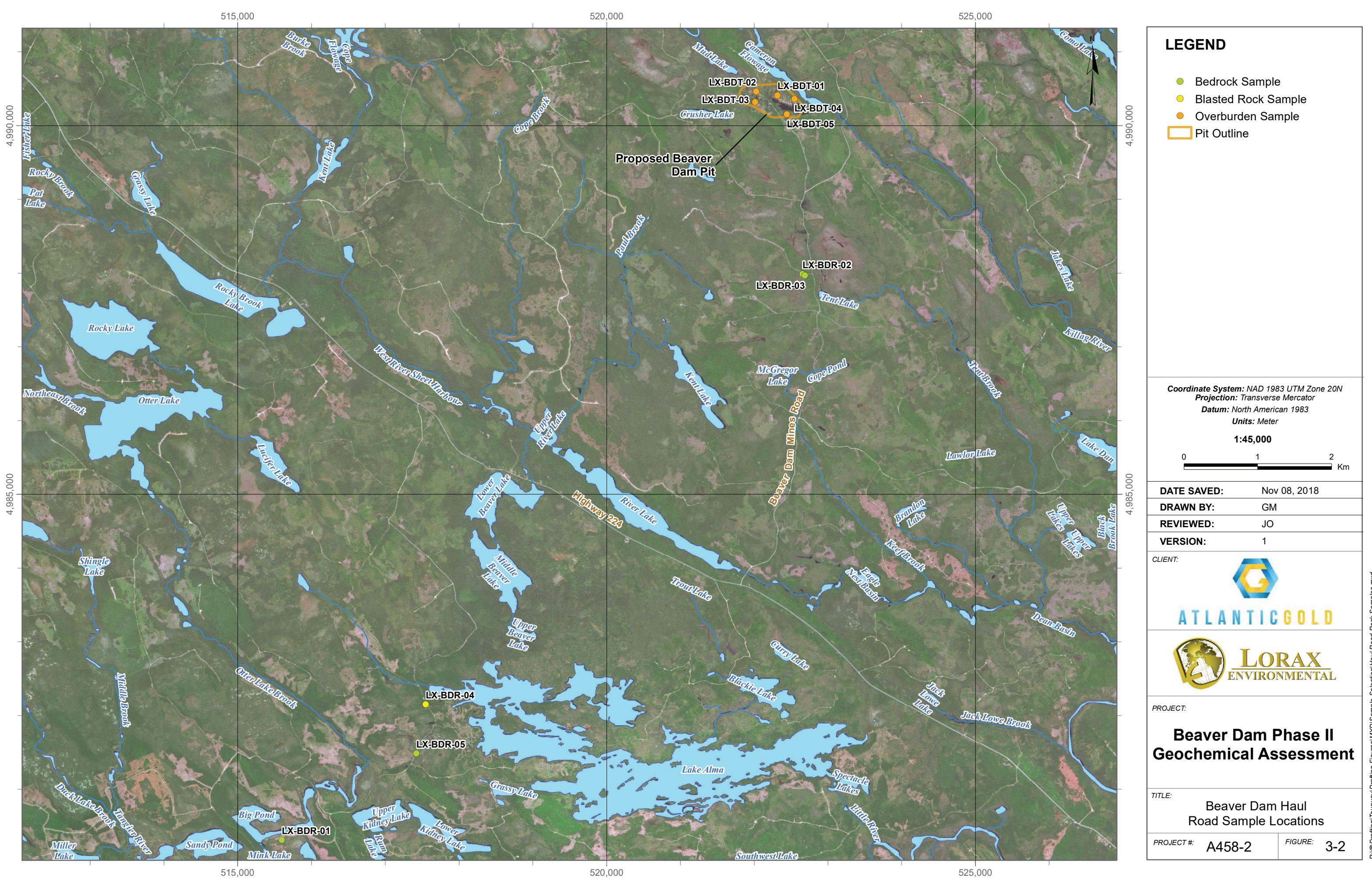
Five rock samples were collected along the Beaver Dam haul road in September 2018 (Figure 3-2 Table 3-4). The geology along the haul road was fairly consistent and all samples collected were greywacke. Samples were collected from outcrops and blasted rock fragments by breaking off rock chips using a rock hammer. These samples were analyzed for ABA and solid phase elements.

## 3.2 Static Test Methods

Static tests carried out to characterize rock and overburden samples include acid base accounting (ABA), aqua-regia digestible elemental abundance, mineralogical investigations, and shake flask extraction (SFE). The SFE tests were only conducted on the five overburden samples. The following sections provide a brief overview of these methods. Static testing was conducted at ALS Laboratories. All analytical procedures were conducted following standard protocols.

### 3.2.1 Acid-Base Accounting

Acid-base accounting (ABA) consists of a series of static tests (paste pH, sulphur species, neutralization potential and acid potential) which are used to evaluate the acid rock drainage (ARD) potential of materials. As materials undergo weathering, the competing influences of acid-generating and alkalinity-producing reactions will determine whether ARD will result. Acidic drainage at mine sites is typically generated from the oxidation of sulphide minerals, whereas neutralization is typically provided by the dissolution of carbonate minerals. The sulphide sulphur content is estimated by the difference between the total sulphur and sulphate sulphur and is used to derive the acid potential (AP) of site materials. The modified neutralization potential (NP) is used to represent the NP of the materials for this site. It is a bulk NP measurement that considers NP from other minerals (*e.g.*, the aluminosilicates) in addition to carbonates. The carbonate NP (CaNP) is calculated using the assumption that the total inorganic carbon (TIC) content is present as calcium carbonate. The relative amounts of NP and AP of a sample can be used to evaluate the potential for acid generation giving consideration to standard regulatory criteria classifying mine solid waste as either PAG (potentially acid generating) or NAG (non acid generating). Consistent with the criteria proposed in Price (2009) to evaluate the likelihood of ARD, materials with a net potential ratio (NPR = NP/AP) less than 2 are classified as PAG, while samples with an NPR value greater than 2 are designated as NAG.



**Table 3-4:**  
**Summary of haul road samples**

Sample ID	Lithology	Sample Type
LX-BDR-01	Greywacke	Bedrock from outcrop
LX-BDR-02	Greywacke	Bedrock from existing quarry
LX-BDR-03	Greywacke	Bedrock from existing quarry
LX-BDR-04	Greywacke	Blasted rock from old quarry
LX-BDR-05	Greywacke	Bedrock from outcrop

### **3.2.2 Aqua-Regia Digestible Elemental Abundance**

Solid phase elemental abundance analyses are conducted on pulverized samples by digesting a sample in aqua regia. The extract is then diluted and analyzed for metals by inductively coupled plasma mass spectrometry (ICP-MS). The data are used to characterize materials and to identify elements of potential environmental concern. The degree of enrichment as compared to average upper continental crust concentrations (AUCCC, Rudnick and Gao, 2014) can provide a general indication of the overall metal enrichment. However, enrichment does not necessarily indicate that the element will become problematic, since the leaching rate is highly dependent on other factors, including the metal-phase associations, grain size, the geochemistry of infiltrating waters and the depositional environment. The solid phase elemental results are also considered when selecting humidity cell samples in order to ensure that representative material is selected to represent the material being characterized by the humidity cell test.

### **3.2.3 Shake Flask Extractions**

Metal contents measured in shake flask extractions (SFE) provide a measure of the mass of readily soluble metals which will be immediately available for leaching upon exposure to infiltrating water. The procedure consists of agitating a representative sample in water, typically at a water to solids ratio of 3:1, for 24 hours. The leachate chemistry can be used as a cursory tool in determining the potential leachate chemistry of water in contact with disturbed rock.

## **3.3 Mineralogy**

Mineralogical analyses are useful in determining the significant forms of acid producing minerals (*i.e.*, sulphides) and acid neutralizing minerals (*i.e.*, carbonates and silicates) in a sample. X-ray Diffraction (XRD) with Rietveld-refinement is a standard technique which provides quantitative mineralogical information. All eight humidity cell subsamples were submitted to the Electron Microbeam & X-Ray Diffraction Facility at the University of British Columbia for mineralogical assessment. Thin sections of the same samples were also investigated by petrographic microscopy

using a Nikon Optiphot polarizing microscope with transmitted and reflected light capabilities. Photomicrographs were taken using a Nikon EOS 70D camera.

### **3.4 Particle Size Distribution**

Particle size distribution (PSD) analyses were carried out to quantify the relative distribution of grain sizes of material placed in the humidity cell tests. It is important to have an estimate of the specific surface area of the humidity cell samples as the kinetic rate at which a material will react is in part dependent on the specific surface area of the sample. The specific surface area is also required when scaling between laboratory and field conditions. Standard mechanical sieving methods were used to characterize the general PSD of samples.

### **3.5 Kinetic Test Methods**

Kinetic testing is used to mimic the natural weathering processes that act on crushed rock or tailings material. The results are used as the basis to predict geochemical loading rates from these materials when stored in surface facilities under oxidizing conditions. Laboratory-based humidity cells are typically composed of a plexiglass cylinder filled with approximately 1 kg of sample crushed to 80% passing <6.4 mm. Beaver Dam materials used for kinetic testing had already been crushed to 80% passing < 2 mm, such that this reduced particle size was used for this test program. The contents of the cells are subjected to moist air for three days, followed by dry air for three days (< 10% relative humidity). At the end of each wet/dry cycle, the contents of the cell are leached with 500 mL distilled de-ionized water on the seventh day (Price, 1997; Lapakko, 2003). The purpose of the leaching step is to recover any readily soluble products that have formed due to mineral dissolution or sulphide oxidation in order to determine the dissolved load contributed from the previous week's test. The leachate is then analyzed for pH, alkalinity and any solutes of interest. Humidity cell tests are usually run for a minimum of 20 cycles (weeks), but it is usually more desirable to continue testing until the rates of sulphate generation and metal leaching have stabilized, which can take up to 120 weeks or more (Lapakko, 2003). These experiments provide data on the primary weathering rates of waste materials and, therefore, the results from this type of testing may be used to estimate the rate of acid generation and metal release to the environment on a weekly basis. As well, after geochemical modeling and careful consideration, these data may be used to estimate drainage chemistry.

## ***4. Results***

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ATLANTIC GOLD

# **4 Results**

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The Beaver Dam geochemical test results are herein presented for three different sets of samples: 1) mine rock; 2) overburden; and 3) haul road material. Static test analyses were carried out on all samples to determine the geochemical characteristics of materials to be disturbed during development of the proposed Beaver Dam Mine. Kinetic tests were carried out on a subset of mine rock samples to predict geochemical behaviour as weathering proceeds. The geochemical results presented have been used to develop source terms for a site-wide water quality model, which is presented under separate cover (Lorax, 2018).

## **4.1 Mine Rock**

### **4.1.1 Static Test Results**

#### **4.1.1.1 Mineralogy**

Mineralogical investigations as part of the ML/ARD characterization program helps provide information on the identity of acid-producing and acid-neutralizing minerals to help interpret static test results and provide a better foundation on which predictions may be based.

Results of the XRD analyses for mine rock samples are summarized in Table 4-1, with the full report results provided in Appendix 4-1. The mineralogical investigation determined that samples are composed primarily of quartz, plagioclase, illite-muscovite, biotite and chlorite with lesser amounts of calcite, pyrite and ilmenite. It is noteworthy that pyrite is the only sulphide phase detected in the Beaver Dam samples by XRD analysis and only detected in 6 of the 8 samples. Calcite is the dominant carbonate mineral in Beaver Dam mine rock, comprising up to 1.9 %. Additional neutralization capacity may be afforded by silicate phases, including plagioclase, chlorite, and biotite, which are present in all samples.

Petrographic analysis of the same eight humidity cell subsamples sheds light on textural relationships within Beaver Dam waste rock and allows the identification of trace minerals that may fall below the detection limit of XRD analysis. Since it is known from site observations and the review of drill hole logs that the greywacke and argillite end-members can be finely-interbedded and occur along a continuum of grain sizes with both material types being represented in all samples, a clear textural distinction had to be made for the purpose of this description. Upon detailed microscopical review, any rock fragments (clasts) within the thin sections that contain primary (*i.e.*, not formed by post-depositional hydrothermal processes) sediment grains of >0.07 mm length or diameter are herein defined as greywacke clasts. For a waste rock fragment to classify as argillite, all primary minerals in a given clast must fall below this threshold. Note that a threshold grain size of 0.07 mm also roughly corresponds with the transition from the silt to sand particle size.

An overview of the petrographic results is shown in Table 4-2 and illustrates that samples HC3 and HC7 have the highest relative proportions of argillite and greywacke clasts, respectively. This breakdown also illustrates that the geologic unit assigned to the eight subsamples may not coincide with the relative greywacke to argillite proportions observed on a thin section scale. In the following, textural relationships will be discussed across the various subsamples specific unless observations were made in particular samples.

Both end-member material types show a similar mineral inventory, although the relative mineral abundances vary. As shown by XRD analysis and confirmed by optical methods, argillite-rich HC subsamples generally contain higher relative amounts of muscovite/illite and chlorite, while showing smaller quartz and feldspar contents. The occurrence of muscovite/illite, chlorite and biotite as matrix (clay) replacement phases in both greywacke and argillite suggests low prograde metamorphism, potentially related to the ore genesis phase.

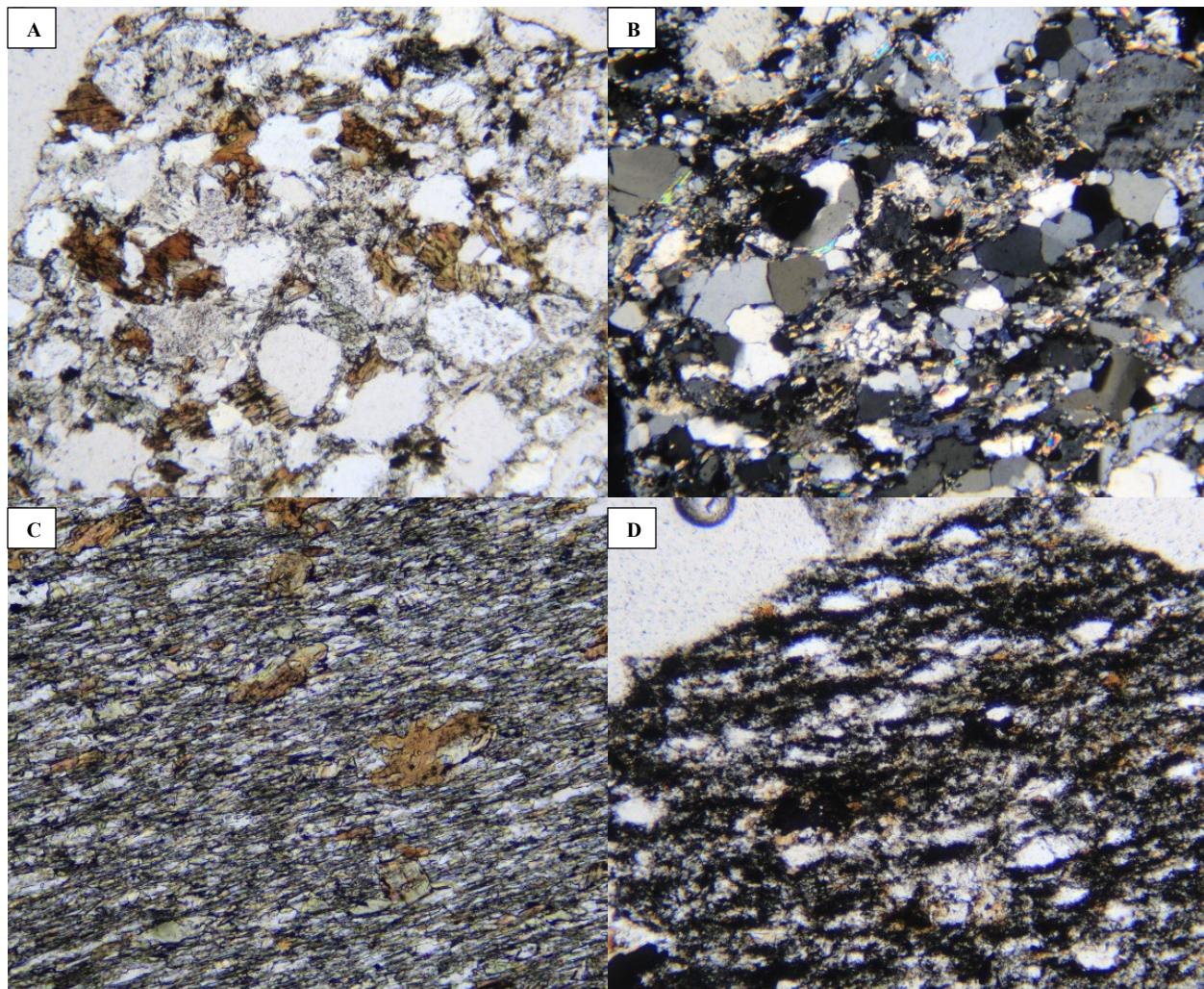
Greywacke fragments are moderately- to well-sorted and primarily composed of rounded quartz and feldspar grains making up >70% of the volume of most fragments. The groundmass (matrix) is typically composed of fine grained sericite, clay, and chlorite filling interstices (Figure 4-1A). Biotite commonly occurs as a medium- to coarse-grained (0.1 – 0.5 mm) hydrothermal alteration phase that may be anhedral or subhedral, although the anhedral variety of biotite is more common in greywacke material. Generally, when present as anhedral patches, biotite typically shows irregular grain boundaries. Grain orientation/foliation is noticeable in the moderately-sorted greywacke clasts (Figure 4-1B) and minor or absent in the coarser-grained well-sorted occurrences (Figure 4-1A). Individual quartz and feldspar grains typically make up a range in grain sized from 0.75 to 0.2 mm, however individual monomineralic grain fragments can reach > 0.5 mm in length.

Argillite aggregates are characterized by an overall finer grain size distribution with a higher relative content of sericite/muscovite and clay minerals. Feldspar and quartz appear as rounded grains and are moderately- to well-sorted. Due to the larger proportion of clay and mica minerals a higher observed degree of foliation is generally observed in argillite fragments (Figure 4-1C), although more randomly-oriented textures are present as well. In some HC subsamples, most notably in HC6, clasts with strongly foliated clay-rich layers (Figure 4-1D) are observed while lacking the typical mica-chlorite matrix replacement pattern typically seen in other argillite fragments. This suggests that portions of the argillite may have undergone less intense hydrothermal alteration than others since the originally deposited clay-rich groundmass is unaltered. Another common textural feature is the retrograde clay-alteration of argillite clasts giving the particle surface a turbid appearance that apparently overprints any prograde mica-alteration patterns. Biotite in argillite samples is commonly distinctly larger-grained than the surrounding groundmass (Figure 4-1C) and forms tabular grains that seem to have grown independently of the sedimentary orientation of within the matrix. Distinct clusters of biotite grains in many are associated with and seem to emanate from secondary ilmenite and magnetite ( $\pm$ rutile,

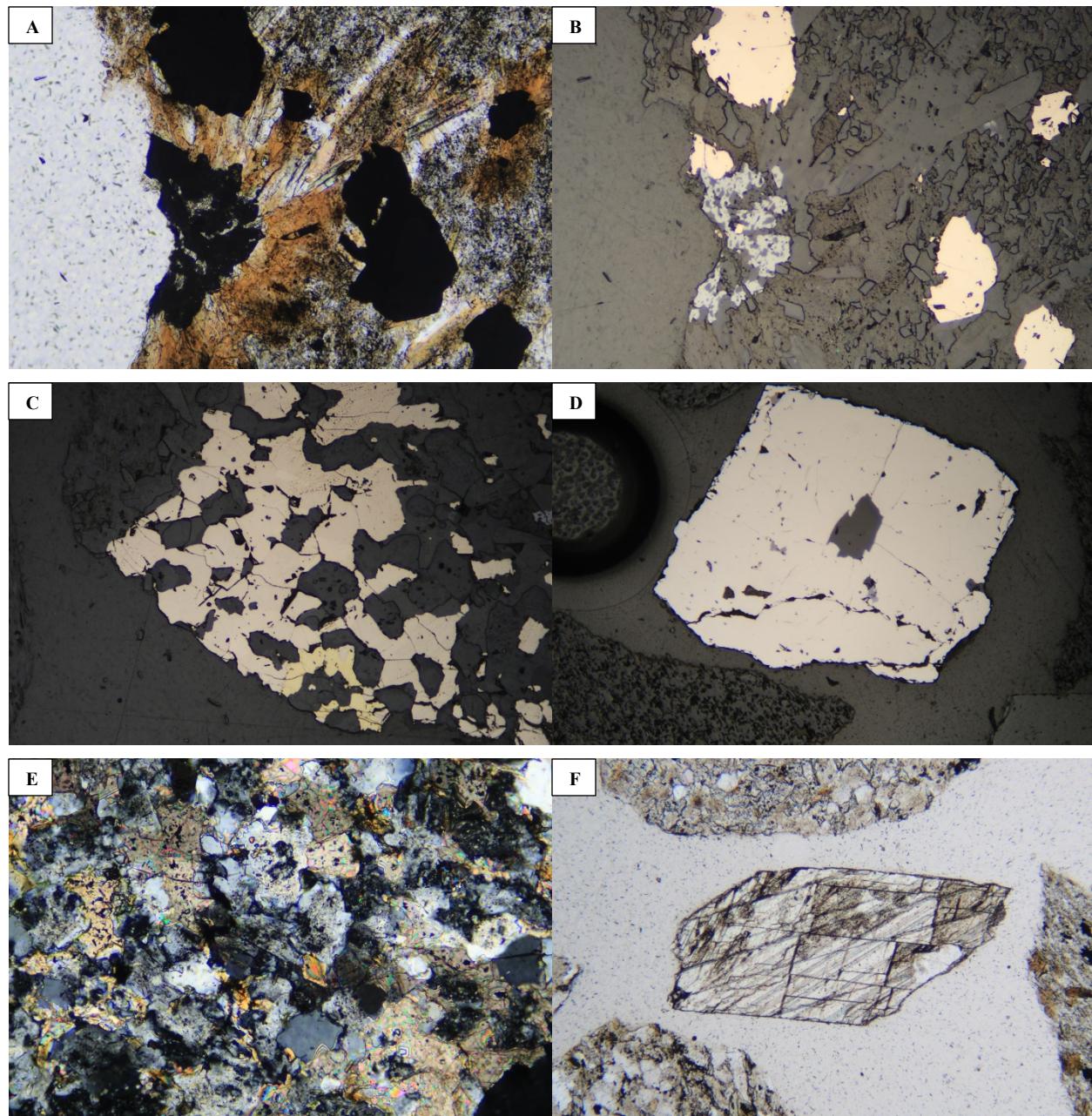
$\pm$ hematite) grains (Figure 4-2 A,B) which make up a significant portion (>50%) of the opaque mineral inventory in most samples and generally display spongy texture.

Sulphides are primarily represented by pyrrhotite (>90%) with minor arsenopyrite, pyrite, and chalcopyrite being observed as accessory sulphide phases. Pyrrhotite occurs as dispersed an- to subhedral grains that may be elongated along the foliation of argillite clasts or as randomly-oriented replacement phases of primary minerals and matrix (Figure 4-2C). Discrete arsenopyrite was only identified in HC8 where, in one instance, it formed a coarse (1.2 mm) subhedral crystal (Figure 4-2D). Chalcopyrite was exclusively identified as minute exsolution blebs in contact with pyrrhotite. Minor pyrite appears to occur primarily as vein fill with only few occurrences as dispersed grain in the argillite groundmass. Both pyrite and pyrrhotite commonly exhibit minor weathering rims along grain boundaries and cracks, however, overall, the degree of oxidation can be described as minor. Due to their fine grain-size, sulphide weathering phases could not be conclusively identified by optical methods alone, however it is assumed that a mixture of poorly crystalline or amorphous hydrous ferric oxides (HFO) and clay minerals make up the observed alteration rims. Notably, two relatively coarse (0.25 mm) pyrite grains observed in HC1 were largely (>50%) weathered by a dark-red, semi-opaque mineral that is identified as goethite.

Carbonate phases were identified in subsamples HC5 and HC7 where they may form large euhedral aggregates (Figure 4-2F) or as anhedral hydrothermal replacement phases, preferentially within the matrix of greywacke clasts (Figure 4-2E). Where observed, carbonate tend to be locally concentrated which may have direct effects on its capacity to buffer acid generated by sulphides.



**Figure 4-1:** Photomicrographs from thin sections of Beaver Dam material. (A) Typical examples of a moderately-sorted greywacke clast (plane-polarized light = PPL). Biotite (brown) and muscovite-sericite fill interstices between rounded feldspar and quartz grains (colourless). Minor chlorite is observed as matrix replacement phase (pale green); FOV = 1.5 mm. (B) Greywacke clast (crossed polars = XPL) that shows minor grain orientation; FOV = 0.65 mm. (C) Typical example of a strongly oriented/foliated argillite clast containing hydrothermal biotite; FOV = 0.65 mm. (D) Argillite clast with very fine-grained clay-minerals (instead of muscovite and chlorite) making up the matrix in between quartz and feldspar grains; FOV = 1.5 mm.



**Figure 4-2:** Photomicrographs from thin sections of Beaver Dam material. (A) Spongy ilmenite ± magnetite (opaque, center left) and pyrrhotite (other opaque phases) associated with cluster of coarse hydrothermal biotite (brown sheaves) under PPL; FOV = 0.74 mm. (B) same as (A) but under reflected light. (C) Cluster of twinned, anhedral pyrrhotite aggregates (light brown) and associated chalcopyrite (bright yellow) filling interstices between quartz and feldspar; FOV = 0.74 mm. (D) Subhedral, fresh arsenopyrite grain observed in HC8; FOV = 1.8 mm. (E) Greywacke clast displaying pervasive carbonate (XPL; brown grains with high interference colours) weathering; FOV = 0.74 mm. (F) Individual, euhedral carbonate grain, likely isolated from vein material; FOV = 1.8 mm.

**Table 4-1:**  
**X-Ray Diffraction Results of Subsamples from the Humidity Cell Test Material**

Mineral Phase	Ideal Formula	Argillite		Argillite-Greywacke		Greywacke		Greywacke-Argillite	
		HC 1	HC 2	HC 3	HC 4	HC 5	HC 6	HC 7	HC 8
		Median S	High S	Median S	High S	Median S	High S	Median S	High S
Quartz low	SiO <sub>2</sub>	31	35	27	33	34	37	51	35
Plagioclase	NaAlSi <sub>3</sub> O <sub>8</sub> – CaAl <sub>2</sub> Si <sub>2</sub> O <sub>8</sub>	21	27	18	19	37	19	29	21
Illite/Muscovite 2M1	K <sub>0.65</sub> Al <sub>2.0</sub> Al <sub>0.65</sub> Si <sub>3.35</sub> O <sub>10</sub> (OH) <sub>2</sub> /KAl <sub>2</sub> (AlSi <sub>3</sub> O <sub>10</sub> )(OH) <sub>2</sub>	29	22	33	29	17	29	11	27
Biotite 1M	K(Mg,Fe) <sub>3</sub> (AlSi <sub>3</sub> O <sub>10</sub> )(OH) <sub>2</sub>	9.9	8.8	8.4	11	5.7	6.8	1.2	9.4
Chamosite	(Fe,Al,Mg)6(Si,Al)4O <sub>10</sub> (OH)8	8.2	6.1	12	7.8	4.0	7.6	5.5	7.0
Calcite	CaCO <sub>3</sub>	-	0.28	0.81	0.25	1.7	-	1.9	0.30
Pyrite <sup>a</sup>	FeS <sub>2</sub>	0.14	0.19	0.098	0.48	0.22	0.29	-	-
Ilmenite	Fe <sup>2+</sup> TiO <sub>3</sub>	0.20	0.51	0.45	-	0.22	0.37	-	0.60

**Notes:** A hyphen indicates the phase was not detected

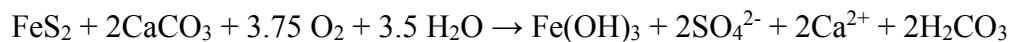
<sup>a</sup>Although pyrite was identified by XRD, the petrographic observations show that pyrrhotite is the more common sulphide mineral

**Table 4-2:**  
**Subsample Summary of Petrographic Observations from the Humidity Cell Test Samples**

Sample ID	Description	Greywacke clasts	Argillite clasts	Sulphides				Carbonates
				Abundance	Type (by abundance)	Weathering		
HC 1	AR - median S	43%	57%	trace (<1%)	Subhedral pyrite (0.05 - 0.25 mm)	Partially replace HFO (>50%), some grains unweathered		N/A
HC 2	AR - high S	50%	50%	trace (<1%)	Anhedral pyrrhotite & pyrite (0.1 - 0.3 mm)	Commonly clay and HFO-weathered rims		N/A
HC 3	AG - median S	20%	80%	1-2%	Anhedral pyrrhotite & pyrite (0.1 - 1 mm)	Mostly unweathered, some grains partially weathered to clay, HFO		N/A
HC 4	AG - high S	42%	58%	2-3%	Anhedral pyrrhotite (0.05 - 0.7 mm) & chalcopyrite	Mostly unweathered, some grains partially weathered to clay, HFO		N/A
HC 5	GA - median S	75%	25%	trace (<1%)	Anhedral pyrrhotite (0.05 - 0.15 mm) & chalcopyrite	Commonly clay and HFO-weathered rims		Matrix replacement patches
HC 6	GA - high S	30%	70%	trace (<1%)	Anhedral pyrrhotite & chalcopyrite (0.1 - 0.2 mm)	Mostly unweathered, some grains partially weathered to clay, HFO		N/A
HC 7	GW - median S	100%	0%	trace (<1%)	Anhedral pyrite (0.02 - 0.05 mm) & chalcopyrite	Mostly unweathered, some grains partially weathered to clay, HFO		Matrix replacement patches, large detached grains
HC 8	GW - high S	50%	50%	trace (<1%)	Anhedral pyrrhotite (0.05 - 0.3 mm), subhedral arsenopyrite (0.2 - 1.1 mm) & chalcopyrite	Mostly unweathered, some grains partially weathered to clay, HFO		N/A

#### 4.1.1.2 Acid Base Accounting Results

The main purpose of ABA characterization is to identify acid-producing and acid-neutralizing phases. ARD will only result from the weathering of sulphide-bearing rocks if there is insufficient alkalinity produced to buffer the acidity generated by the sulphide oxidation process. The oxidation of pyrite produces two sources of acid; one from the oxidation of sulphide and the other from the oxidation/hydrolysis of iron. If reactive neutralizing minerals are present, they will not stop the oxidation reactions but may buffer the pH to prevent the development of ARD. While many mineral dissolution reactions can be thought of as acid buffering, the minerals most typically responsible for acid neutralization are fast dissolving carbonates such as calcite and dolomite. At a pH < 6.3, the sulphide oxidation and carbonate neutralization reaction is typically expressed as:



However, it should be noted that not all carbonate minerals neutralize acid as effectively as others. For example, Fe-bearing carbonate minerals, such as ankerite and siderite, are much less effective at neutralizing acid compared to calcite due to the fact that the  $\text{Fe}^{2+}$  liberated is oxidized to  $\text{Fe}^{3+}$ , which then precipitates as  $\text{Fe(OH)}_3$  producing acidity in the process. As a result, the net capacity of a sample to neutralize acid decreases as the amounts of Fe-bearing carbonates increases (Jambor *et al.*, 2003). Silicate minerals may also contribute to the total neutralizing capacity of a sample; however, rates of silicate dissolution are much slower and thus may limit the ability of these minerals to buffer acid generation.

The full set of ABA analyses for all mine rock samples is presented in Appendix 4-2 and includes paste pH, total S, sulphate S, sulphide S, acid potential (AP), total inorganic C, total C, modified neutralization potential (NP) and fizz rating. A summary of the mine rock ABA results is provided in Table 4-3, with the humidity cell test subsample results presented in Table 4-4.

##### 4.1.1.2.1 Paste pH

Paste pH provides an indication of whether a sample is currently generating acidity at the time of sampling. Paste pH values for the all samples range from 7.9 to 9.5, indicating that these samples are not currently acid generating. The GA samples show the widest range of pH values, ranging between 7.9 and 9.5, with a median pH value of 9.0 (Table 4-3).

**Table 4-3:**  
**Summary of acid-base accounting results for mine rock lithologies**

<b>Sample ID</b>	<b>Paste pH</b>	<b>Total S</b>	<b>Sulphate S</b>	<b>Sulphide S</b>	<b>Total C</b>	<b>CaNP</b>	<b>Modified NP</b>	<b>NPR</b>
		%	%	%	%	kg CaCO <sub>3</sub> /t	kg CaCO <sub>3</sub> /t	ModNP/AP
<b>ARGILLITE (AR) n = 10</b>								
Min	8.2	0.020	0.010	0.020	<0.05	<4.5	6.0	0.27
Median	9.0	0.44	0.020	0.44	<0.05	<4.5	8.5	0.64
Max	9.5	1.0	0.060	0.94	<0.05	<4.5	37	17
<b>ARGILLITE-GREYWACKE (AG) n = 11</b>								
Min	7.9	0.010	<0.01	0.010	<0.05	<4.5	7.0	0.23
Median	9.1	0.12	0.020	0.12	<0.05	<4.5	8.0	8.0
Max	9.5	1.1	0.060	1.1	0.17	14	177	47
<b>GREYWACKE-ARGILLITE (GA) n = 14</b>								
Min	7.9	0.010	0.010	0.010	<0.05	<4.5	5.0	0.40
Median	9.0	0.16	0.020	0.14	0.11	9.1	9.0	2.5
Max	9.5	1.7	0.060	1.5	0.57	48	58	46
<b>GREYWACKE (GW) n = 24</b>								
Min	8.9	<0.01	<0.01	<0.01	<0.05	<4.5	6.0	0.16
Median	9.2	0.025	0.020	0.030	0.10	9.1	14	22
Max	9.5	1.2	0.060	1.2	0.60	50	160	512
<b>QUARTZ VEIN (QTZV) n = 3</b>								
BD157-070	9.2	0.160	0.020	0.140	-	-	145	33
BD156-108	9.3	0.060	0.010	0.050	-	-	10	6.4
BD160-089	8.9	0.040	0.010	0.040	-	-	30	24

**Notes:** n = number of samples used in statistical distribution

A hyphen indicates the parameter was not analyzed.

Values in grey italics are below the analytical detection limit. Values were set at the detection limit for calculation of NP, AP and NPR values.

Sulphate S is calculated using the HCl method.

AP (acid potential) calculated using sulphide sulphur (% non-sulphate sulphur x 31.25);

CaNP (carbonate neutralization potential) calculated using total inorganic carbon (% TIC x (100.09/12.01) x 10);

Modified NP is obtained by the modified Sobek method.

NPR = neutralization potential ratio; calculated as Modified NP / AP.

**Table 4-4:**  
**Summary of acid-base accounting results for humidity cell subsamples**

<b>HC ID</b>	<b>Description</b>	<b>Paste pH</b>	<b>Total S</b>	<b>Sulphate S</b>	<b>Sulphide S</b>	<b>Total C</b>	<b>CaNP</b>	<b>Modified NP</b>	<b>NPR</b>
			%	%	%	%	kg CaCO <sub>3</sub> /t	kg CaCO <sub>3</sub> /t	ModNP/AP
<b>ARGILLITE (AR)</b>									
HC1	Median-S	8.9	0.02	0.02	0.02	<0.05	<4.5	6	9.6
HC2 <sup>a</sup>	High-S	8.7	0.25	0.015	0.25	<0.05	<4.5	7.5	1.1
<b>ARGILLITE-GREYWACKE (AG)</b>									
HC3 <sup>a</sup>	Median-S	8.7	0.20	0.02	0.18	0.11	9.1	15	5.6
HC4 <sup>a</sup>	High-S	8.2	0.77	0.015	0.76	<0.05	4.5	8.0	0.45
<b>GREYWACKE-ARGILLITE (GA)</b>									
HC5 <sup>a</sup>	Median-S	8.5	0.21	0.015	0.19	0.18	15	22	11
HC6 <sup>a</sup>	High-S	8.6	0.39	0.015	0.39	<0.05	<4.5	5.5	0.47
<b>GREYWACKE (GW)</b>									
HC7	Median-S	9.1	0.03	<0.01	0.03	0.2	15.9	23	25
HC8 <sup>a</sup>	High-S	9.3	0.36	0.03	0.35	0.05	4.5	8.0	0.74

**Notes:**

<sup>a</sup>Humidity cell is made up of two samples. An average value is presented and a 1:1 mixture is assumed.

Values in grey italics are below the analytical detection limit. Values were set at the detection limit for calculation of NP, AP and NPR values.

Sulphate S is calculated using the HCl method.

AP (acid potential) calculated using sulphide sulphur (% non-sulphate sulphur x 31.25);

CaNP (carbonate neutralization potential) calculated using total inorganic carbon (% TIC x (100.09/12.01) x 10);

Modified NP is obtained by the modified Sobek method.

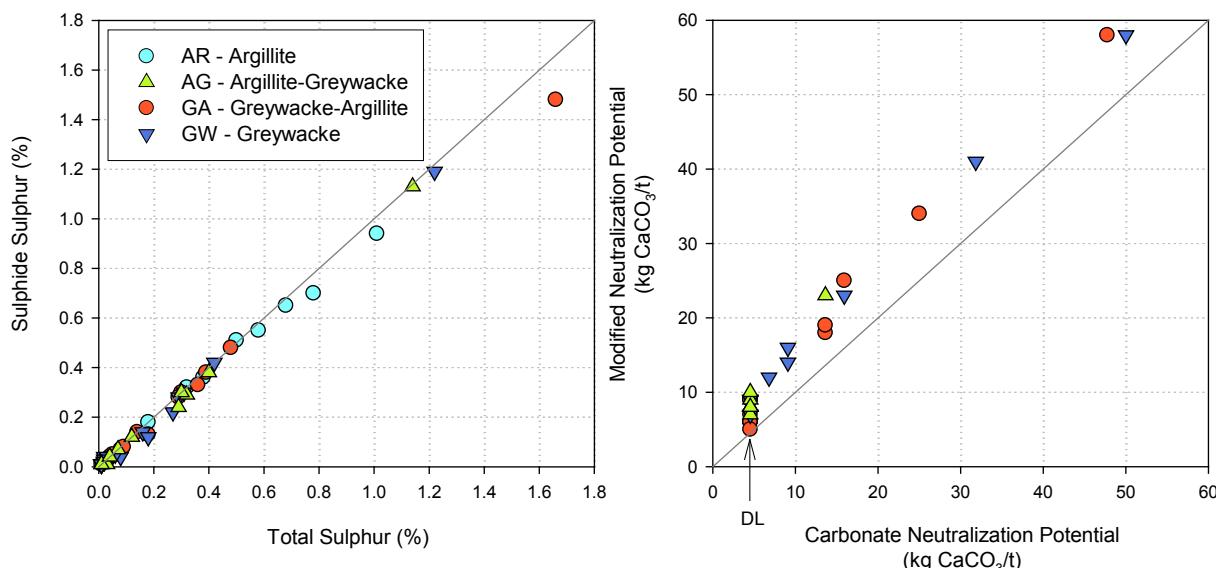
NPR = neutralization potential ratio; calculated as Modified NP / AP

#### 4.1.1.2.2 Sulphur Species and Acid Potential

The total sulphur (T-S) contents of the Beaver Dam samples vary from below detection (0.01%) to 1.7% (Table 4-3). Of the four main rock types, the AR samples have the highest total median sulphur content (0.44%) and the GW samples have the lowest (0.025%). The three quartz vein samples analysed had low sulphur contents with a mean value of 0.087%. The sulphate S contents are generally low in most samples and typically fall below detection (0.01%) but can reach up to 0.06%. The sulphide S contents are strongly correlated with T-S (Figure 4-3) and indicate that most of the sulphur present in samples is in the form of sulphide. The sulphide S contents range from less than detection level values (0.01%) up to a maximum value of 1.5%, as measured in a GA sample (Figure 4-3). The AR samples have the highest median sulphide S contents (0.44%), whereas the GW samples have the lowest median sulphide S contents (0.030%).

The mineralogy results indicate that the sulphide S is primarily in the form of pyrrhotite, which is acid generating. The acid potential (AP) of samples is calculated based on the sulphide S values and used in the determination of the neutralization potential ratio (NP/AP) discussed in Section 4.1.1.2.4.

In general, the humidity cell test subsamples are considered to be representative of median and high S contents for the different lithologies; however, it is noted that the HC1 and HC2 have low sulphide S contents as compared to the median AR value (0.44%). These cells also have low NP contents and were specifically composited to generate acidic leachates in a timely manner.



**Figure 4-3:** Plots showing total sulphur *versus* sulphide sulphur and carbonate neutralization potential *versus* modified neutralization potential for the Beaver Dam mine rock samples.

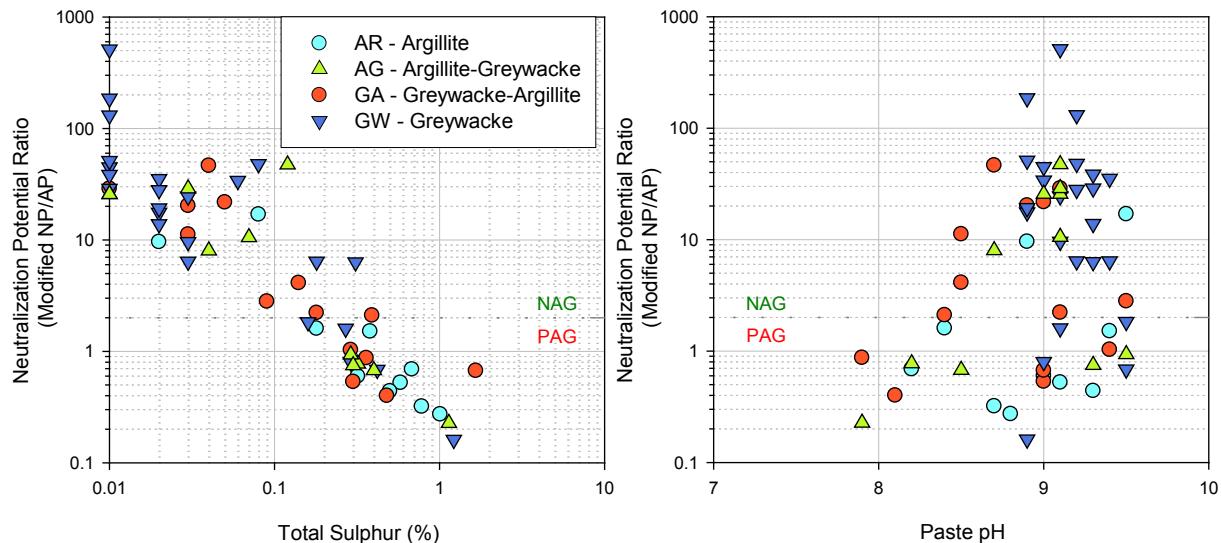
#### 4.1.1.2.3 Neutralization Potential

The total inorganic carbon (TIC) content of Beaver Dam samples ranges from detection level values (<0.05%) up to 0.60%, with the highest median values measured in the GW and GA samples (0.10 % and 0.11%, respectively; Table 4-3). It is assumed that the inorganic C is present as carbonate minerals and thus TIC values are used to calculate the carbonate neutralization potential (CaNP). The CaNP values of Beaver Dam samples range from detection level values (<4.5 kg CaCO<sub>3</sub>/t) up to 50 kg CaCO<sub>3</sub>/t. The GW and GA samples have the highest median CaNP values at 9.1 kg CaCO<sub>3</sub>/t, while the AR samples have the lowest median CaNP values (< 4.5 kg CaCO<sub>3</sub>/t; Table 4-3).

The CaNP values are generally comparable to the modified NP values (Figure 4-3), although the modified NP values are consistently higher suggesting that minerals other than carbonate (*i.e.*, aluminosilicates) are available for acid-neutralization. This NP is not as readily available as the carbonate NP; however, when sulphide values are low, as is the case at Beaver Dam, the NP from the dissolution of non-carbonate minerals will contribute to the neutralization of the low rates of acid production. Silicate minerals that act as neutralizing agents may include plagioclase, biotite, chlorite, and certain clay minerals, all of which were identified in the mineralogical analysis. For the remainder of this assessment, modified NP is used as the basis for NPR (NP/AP) calculations in order to quantify a material's ARD potential.

#### 4.1.1.2.4 Neutralization potential ratio

The neutralization potential ratio (NPR) is calculated as the ratio of NP to AP and is presented based on the ModNP (Table 4-3; Figure 4-4). In the absence of long-term kinetic test data, the NPR value is the most important parameter in the evaluation of a material's likelihood to generate ARD. In accordance with guidelines set out in Price (2009), a value >2 is considered to be NAG, whereas a value <2 is considered to be PAG. Based on the results, a significant number of samples covering all four major lithologies are PAG (Table 4-5).



**Figure 4-4:** Plots showing neutralization potential ratio (NPR) versus total sulphur for the Beaver Dam mine rock samples, showing the potentially acid generating (PAG) boundaries (dark grey dashed line). Plot (A) presents NPR values as modified neutralization potential (NP) versus total sulphur. Plot (B) presents NPR values as Modified NP versus paste pH. Lithology codes defined in Table 3-1.

**Table 4-5:**  
**PAG Proportions for the Beaver Dam mine rock**

Lithology	% PAG	% NAG
Argillite (AR)	80	20
Argillite-Greywacke (AG)	45.5	54.5
Greywacke-Argillite (GA)	35.7	64.3
Greywacke (GW)	20.8	79.2
<b>Total</b>	<b>39.0</b>	<b>61.0</b>

#### 4.1.1.3 Total Solid Phase Elemental Analysis Results

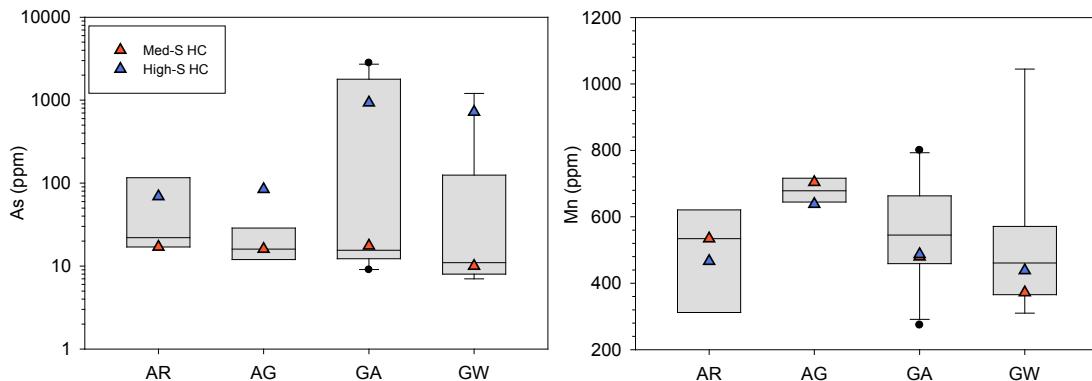
The results of the total solid phase elemental analysis are presented in Appendix 4-3 and summarized in Table 4-6. Elements that are greater than 3× their respective AUCCC values (Rudnick and Gao, 2014) and greater than 2x the detection limit are listed in Table 4-6 and include Ag, As, Cu, Mn, and Pb. Solid phase concentrations of Ag, As, Mn, and Pb are enriched by a factor greater than 10 above the AUCCC in one or more samples. It is noteworthy that the elevated As concentrations (>10× AUCCC) occur in all Beaver Dam lithologies. Elevated Mn values (>10× AUCCC) occur in three of the 4 lithologies, whereas elevated Ag and Pb (>10× AUCCC)

are limited to the AG samples. The AUCCC for Ag (0.053 ppm) is considerably lower than the analytical detection limit of 0.2 ppm so any values above or equal to the detection limit are flagged by the 3x the AUCCC screening criteria. Only two samples in the dataset have Ag values above 2x the detection limit.

A summary of the humidity cell total solid phase elemental results is included in Table 4-7. The concentrations of key metals of potential concern (As and Mn) in humidity cell subsamples are presented as compared to the range for each of the major rock types (Figure 4-5). Arsenic specifically is highest in HC 6 and HC 8.

The elemental enrichments highlight elements that require additional scrutiny in leaching tests. However, an element present at an elevated concentration in the solid phase may not become a metal leaching issue and vice versa. There are several factors that influence the leaching rates of elements, including the mineral association and stability, as well as the chemistry of the water in contact with the rocks. Kinetic test results provide a better indication of the leaching potential.

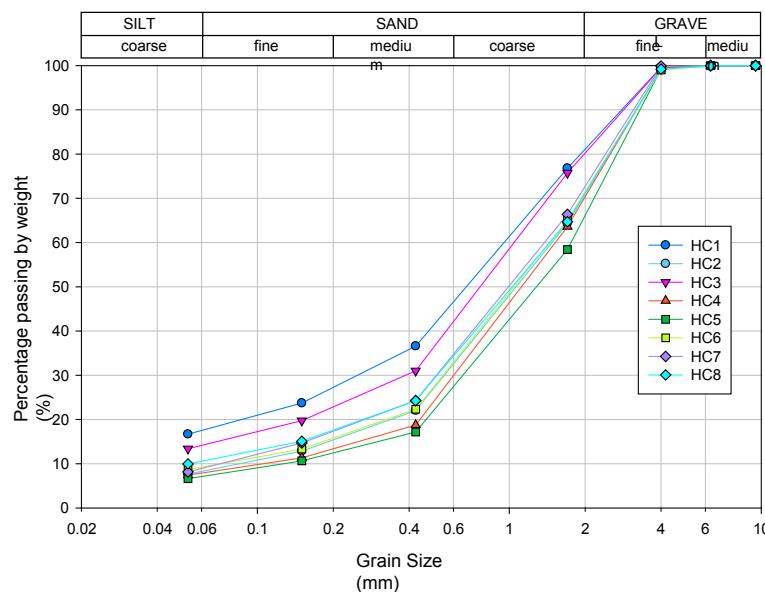




**Figure 4-5:** Box plots showing the range of As and Mn concentrations as compared to median and high sulphur humidity cell subsamples. Note that the box limits represent 25<sup>th</sup> and 75<sup>th</sup> percentile levels, the horizontal bar represents the median level and the whiskers represent the 5<sup>th</sup> and 95<sup>th</sup> percentile values.

#### 4.1.1.4 Particle Size Distribution Results

The particle size distribution results for the humidity cell subsamples are presented in Appendix 4-4 and summarized in Figure 4-6. All humidity cell samples were mechanically crushed to  $P_{80} < 2$  mm by the laboratory prior to placement in the humidity cells. This was done erroneously as the standard particle size to be placed into humidity cells should be  $P_{80} < 6.4$  mm (Price, 2009). The finer particle size will be accounted for in the use of humidity cell data for source term development. The PSD results show that samples HC1 (AR) and HC3 (GA) have relatively higher proportions of fines as compared to other humidity cell samples.



**Figure 4-6:** Particle Size Distribution results for humidity cell test samples presented as percent passing by weight versus grain size.

#### 4.1.2 Kinetic Test Results

Laboratory kinetic test procedures are designed to quantify weathering rates under standardized conditions. During the initial cycles of laboratory kinetic testing, sulphate and metals often have highly variable release rates before stabilizing at a relatively constant rate (Sapsford *et al.*, 2009). This variability is a response to exposure of fresh surfaces of crushed samples and the dissolution of stored oxidation products that may have accumulated on the samples during storage prior to being placed in a humidity cell. Once exposed mineral surfaces have equilibrated to this environment, stable reaction rates can be determined. Humidity cells often require many weeks to approach geochemical stability, and reaction rates rarely remain constant on a week-to-week basis. It should be noted that aqueous concentrations in the weekly/biweekly rinse water should not be considered as direct predictions of on-site drainage chemistry due to the high water to solid ratio used in this type of testing (Sapsford *et al.*, 2009). In reality, conditions within large-scale mine rock storage facilities are much different with lower water/rock ratios, incomplete flushing of particle surfaces, and secondary minerals frequently reaching saturation and precipitating out of solution.

The kinetic test program for the Beaver Dam mine rocks consists of eight humidity cells covering the four main lithologies (AR, GA, AG and GW). Samples descriptions are provided in Table 3-2 and experimental methods are described in Section 3.5. Static characterization testwork, including mineral identification, ABA, solid phase element determination and PSD were completed on subsamples of each of the kinetic tests (presented in Section 4.1.1). Results of the static testwork show that the subsamples are considered to be representative of the respective lithologies. This report provides interim results from the ongoing humidity cell tests that were initiated in April 2018 and have been running for 24 weeks (at the time of reporting) with weekly flushing (500 mL) and sampling events. The full set of interim leachate results are presented in Appendix 4-5 and summarized in the sections below.

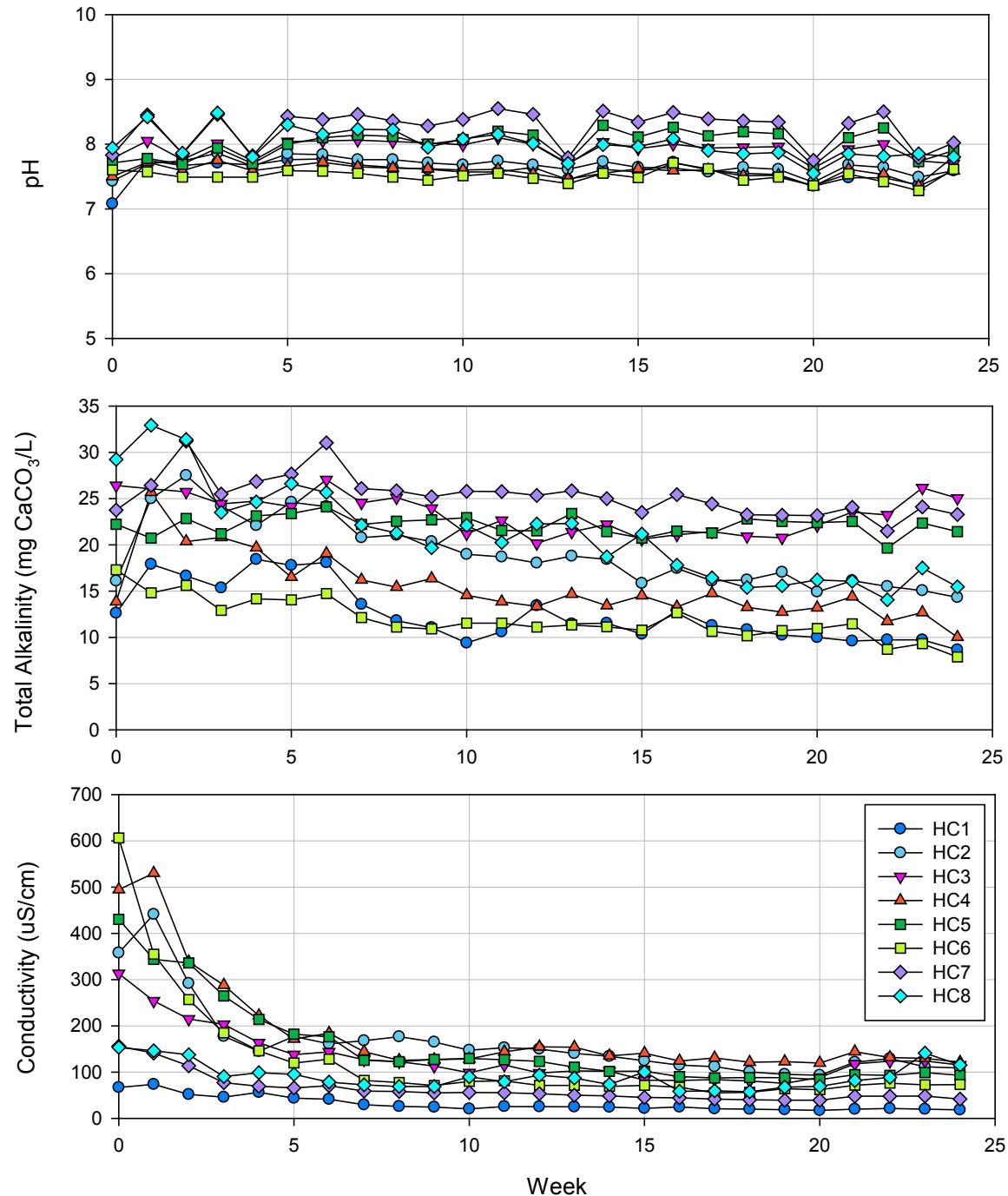
##### 4.1.2.1 pH and Sulphate Loading

The leachate from all humidity cells remains circum-neutral. All pH values are between 7.1 and 8.6 (Figure 4-7). In general, when comparing the pH values for humidity cells composed of the same lithology, the median S humidity cell has a higher pH, relative to the high S humidity cell. The two AR humidity cells are an exception as the pH from HC1 (median S) is slightly lower than the pH from HC2 (high S), although the pH values from these two humidity cells are comparable. HC1 also has lower total alkalinity relative to HC2 (Figure 4-7). Overall, HC7 (median S, GW) has the highest pH values (range: 7.8 to 8.6) and HC6 (high S, GA) has the lowest pH values (range: 7.3 to 7.7). Note that the minimum pH for the dataset (7.1) was in cycle 0 in HC1.

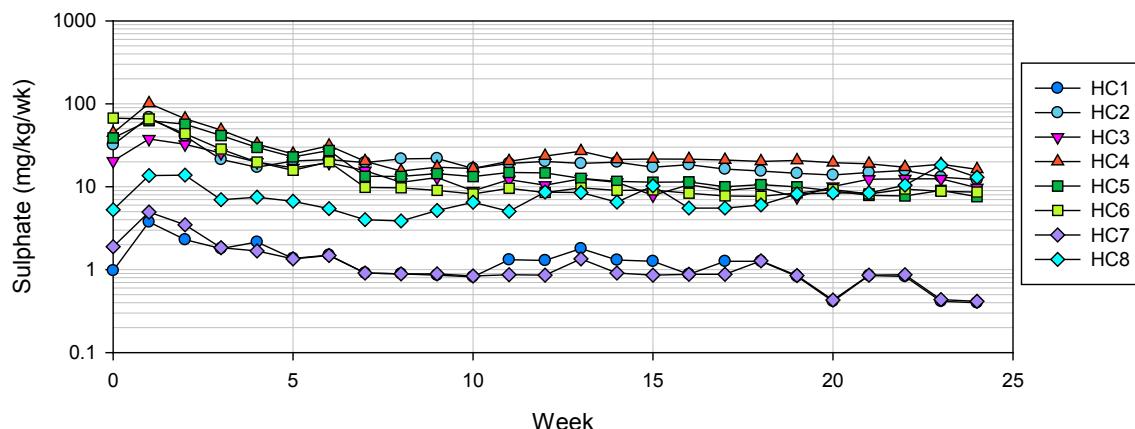
Leachate concentrations from laboratory kinetic tests represent primary sulphate and metal release rates that do not fully account for the influence of field conditions that govern the actual sulphate and metal concentrations in mine drainage. In addition, concentrations measured from humidity cells are susceptible to changes in the volume of water added and collected at the end of each cycle, and hence, concentration data do not provide a strictly quantitative estimate of drainage chemistry. To provide a more functional parameter which can be used to compare results between different humidity cells, sulphate and metal concentrations in mine rock leachate are normalized to the mass of sample in the humidity cell and the volume of leachate collected each week, producing weekly mass loadings (mg solute/kg sample/wk).

For example, the rate of sulphide oxidation in mine rock samples is estimated from sulphate loadings calculated from the humidity cell leachates. To obtain sulphide oxidation rates, weekly leachate sulphate concentrations are multiplied by the volume of leachate collected from each column and cell to obtain a mass loading which is independent of the volume of leachate collected from each humidity cell. These loadings are then normalized to the mass of sample in the humidity cell. This normalized loading calculation is an important step because it allows the direct comparison of leachate data between cycles within the same humidity cell, as well as between different humidity cells.

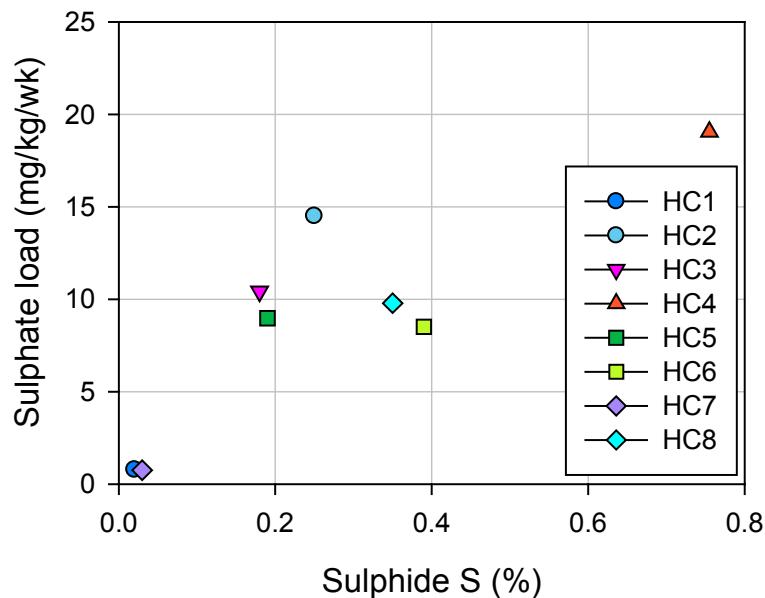
The majority of the humidity cells initially show high sulphate loading rates that decrease and begin to stabilize between weeks 5 and 10 (Figure 4-8). The elevated values in the first few weeks of the tests are due to dissolution and flushing of readily soluble surface oxidation products (*e.g.*, secondary sulphates). HC1 and HC7 have the lowest sulphate loading rates, with loads decreasing to less than 2 mg/kg/wk after week 5. These humidity cells also have the lowest sulphide S (0.02% and 0.03%, respectively). HC2 through HC6 have shown relatively stable or slightly decreasing sulphate loading rates in the later cycles with loading rates between approximately 10 and 20 mg/kg/wk. HC4 has the highest sulphate loading rates with values remaining above 15 mg/kg/wk. Sulphide S is also highest in this humidity cell (0.76%). HC8 initially had a lower sulphate loading rate (<15 mg/kg/wk) but has shown an increasing trend up to 18.3 mg/kg/wk in week 23. Further cycles will determine if this trend continues or if the sulphate loading rates will stabilize. The average sulphate loading rate for the last 8 cycles is plotted with sulphide S content (Figure 4-9). There is some positive correlation between the sulphate loading rate and sulphide S.



**Figure 4-7: pH, total alkalinity, conductivity and sulphate in leachates from Beaver Dam humidity cells**



**Figure 4-8:** Sulphate loading rates in Beaver Dam humidity cell leachates



**Figure 4-9:** Average sulphate loading rate versus sulphide S. Sulphate load is an average from weeks 17-24.

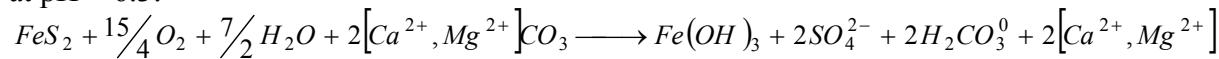
#### 4.1.2.2 Carbonate Molar Ratio

The carbonate molar ratio (CMR) is a proxy for the rate of carbonate dissolution (NP depletion) and sulphide oxidation occurring in the laboratory test reactor, assuming that the base cations are derived only from the NP source and the sulphate is derived from the oxidation of pyrite. The CMR is calculated as:

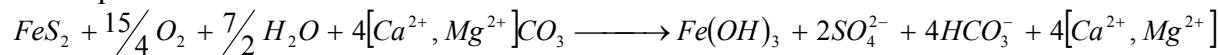
$$CMR = \frac{[Ca^{2+}] + [Mg^{2+}]}{[SO_4^{2-}]}$$

When reactive sulphides are present, the dominant NP consumption reaction is that of acid neutralization. In the most simplistic scenario, when carbonate minerals are present, the oxidation-neutralization reaction is pH-dependent. Assuming no Ca and SO<sub>4</sub> are lost to gypsum precipitation, two carbonate consumption reactions can describe the process, including:

at pH < 6.3:

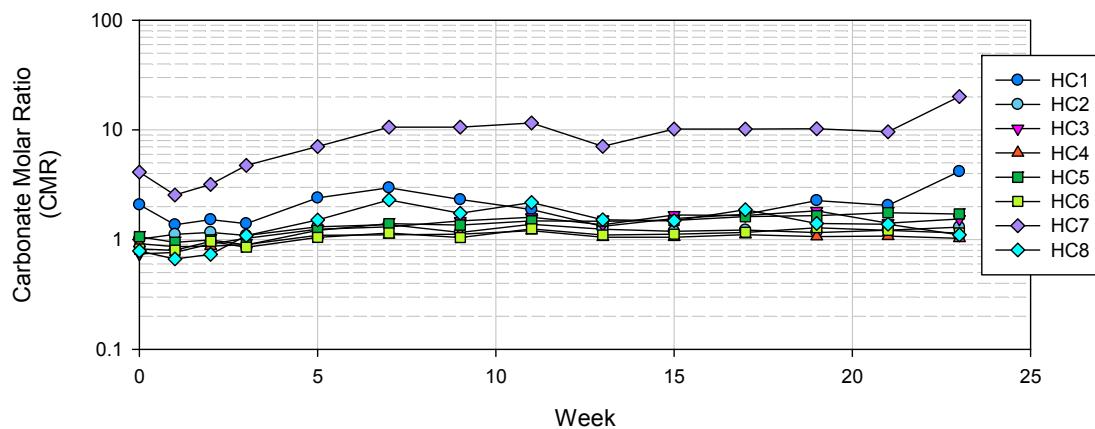


at 6.3 < pH < 10.3:



Neutralization of acidity up to pH levels of 6.3 produce one mole of Ca (+ Mg) for each mole of SO<sub>4</sub> released, producing a CMR = 1.0. At pH levels above 6.3, H<sub>2</sub>CO<sub>3</sub> is not the dominant form of inorganic carbon in an aqueous solution and the bicarbonate ion (HCO<sub>3</sub><sup>-</sup>) is by far the most abundant. Thus, at near-neutral pH levels, calcium carbonate is less efficient at neutralizing acidity and twice as much carbonate is required to produce a balanced solution. Under these conditions, 4 moles of Ca (+ Mg) are theoretically released relative to 2 moles of SO<sub>4</sub> producing a CMR = 2.0. The relationships derived from these chemical equations assume that pyrite oxidation is the sole source of sulphur and iron in the product that take the form of sulphate and iron hydroxide, respectively. Thus, the oxidation of other sulphide minerals, dissolution of soluble sulphate minerals, the formation of other secondary products, or dissolution of carbonates by dilute waters in the absence of significant sulphide oxidation may alter the relationship (Mattson, 2005).

The CMR for the majority of the humidity cells is between 1 and 2 (Figure 4-10). HC1 and HC8 have slightly higher CMR values between 1 and 3. HC7 has the highest CMR values which have increased from approximately 2.5 up to 20.



**Figure 4-10: CMR values for Beaver Dam humidity cell leachate.**

The CMR can be used to calculate the CaNP depletion rate which can in turn be used to calculate the amount of time required to consume all available CaNP from the humidity cell samples. Theoretically, once carbonate is depleted and non-carbonate phases are responsible for buffering the pH, the CMR will equilibrate around a value of 1, which allows for a slower depletion of the remaining NP. For the purpose of this high-level assessment, it was conservatively assumed that the CMR and carbonate depletion rate can also be applied to all NP sources that were quantified by the ModNP method used for the Beaverdam ABA analyses. The NP depletion rate is calculated as follows:

$$\text{NP depletion rate} = \text{CMR} \times \text{Sulphate loading rate (in kg CaCO}_3/\text{t/wk equivalents)}$$

For the PAG humidity cells that are expected to become acidic, the time to NP depletion theoretically corresponds to the time to onset of acid generating conditions, although the grain liberation is an important factor. Note that for the purposes of calculating NP depletion rates and the time for complete depletion of NP, the initial sulphate production rates, which reflect the flushing of non-acid generating surface oxidation products such as gypsum, were not incorporated in the calculation. Rather the relatively stable NP depletion rates in later cycles of the tests more appropriately reflect depletion based on sulphate produced by sulphide oxidation. This will prevent underestimating the amount of available NP and the lag time to the onset of acidic conditions. The model results suggest that NP will be depleted between around 7 to 11 years for the PAG humidity cells. To provide a more realistic estimate that considers the slowing of sulphide oxidation at the Beaver Dam site in response to colder temperatures and/or freezing near-surface rock, the sulphate leaching rate was corrected by a factor of 0.3. This value is consistent with the temperature-dependent sulphide oxidation rate and sulphate release rate observed around 10°C (Dockrey and Mattson, 2016), which is somewhat higher than the mean annual temperature measured at the Beaver Dam site. Note that this approximation is only applicable for waste rock exposed to ambient air temperatures such as pit walls, since waste rock piles may have a different temperature profile. Plotting humidity cell total sulphur ( $S_T$ ) contents against the respective temperature-corrected NP depletion rates shows a positive correlation (Figure 4-11a) that can be written as:

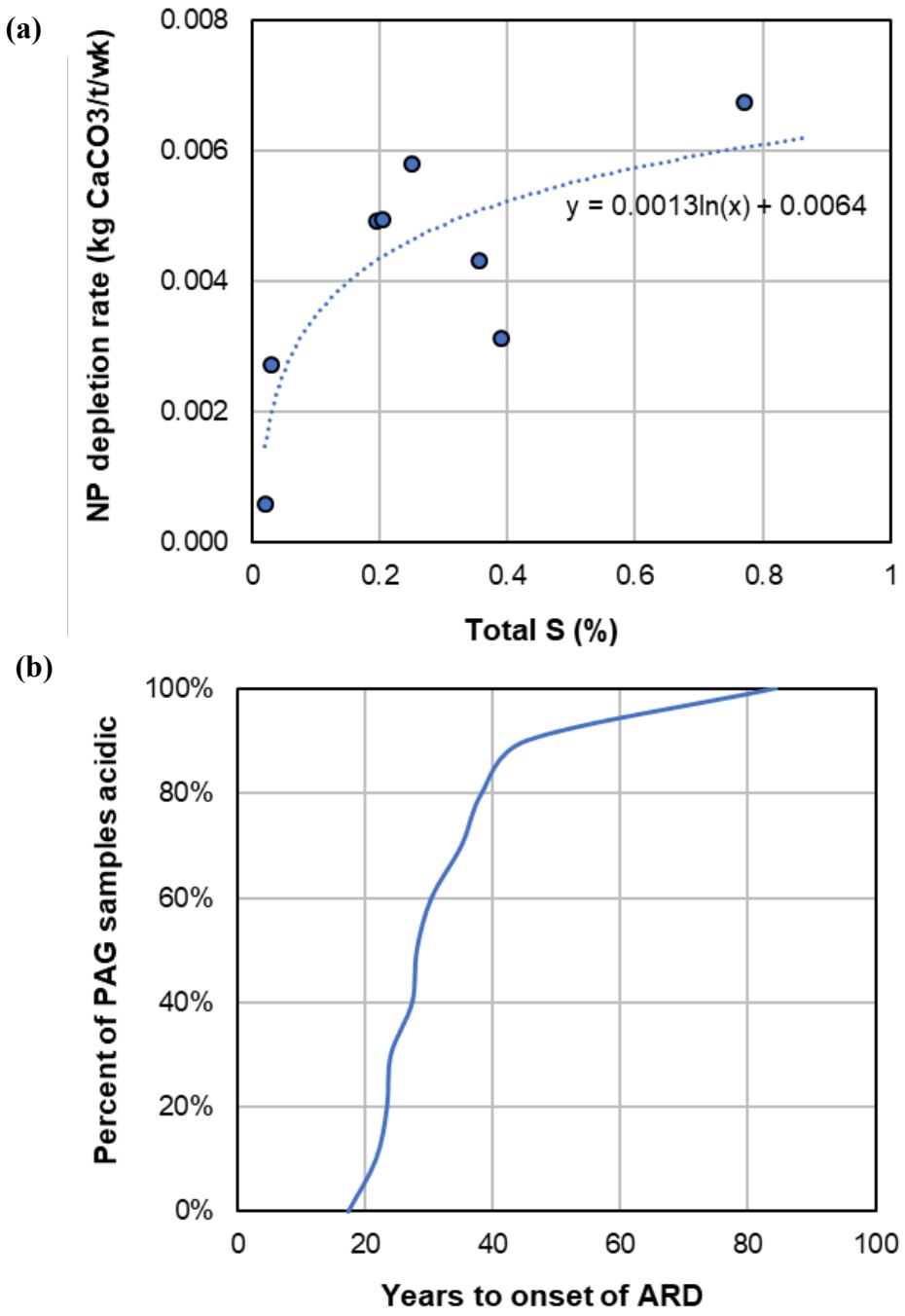
$$\text{NP depletion rate} = 0.0013 \times \ln [S_T] + 0.0064$$

This function was then applied to the PAG samples within the Beaver Dam static test population to quantify a range of lag times that can be expected until these PAG samples release net acidic drainage. The results of this exercise are shown in Figure 4-11b and show that it will take around 20 years for 10% of all PAG samples and 28 years for 50% of all PAG samples to turn acidic.

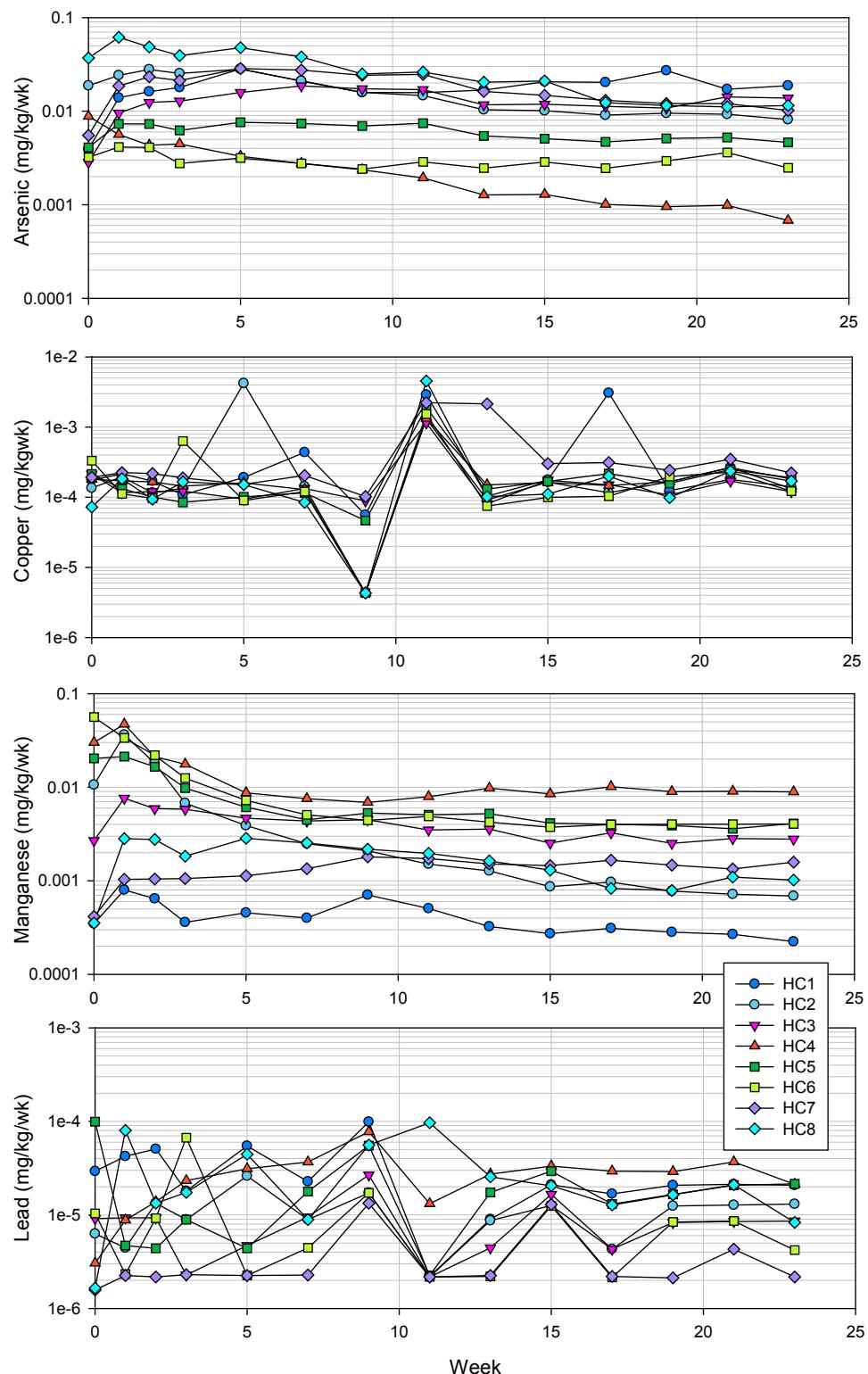
#### 4.1.2.3 Metal Leaching Trends

The trends in leachate mass loading rates over time are provided in Figure 4-12 for selected species. These species have elevated concentrations with respect to 3x the AUCCC in the mine rocks (Section 4.1.1) and include As, Cu, Mn and Pb. Although the Ag concentrations in the total solid phase elemental analysis were above 3 times the AUCCC, Ag is not included in the discussion in this section as all Ag concentrations in the humidity cell leachates remain below the analytical detection limit.

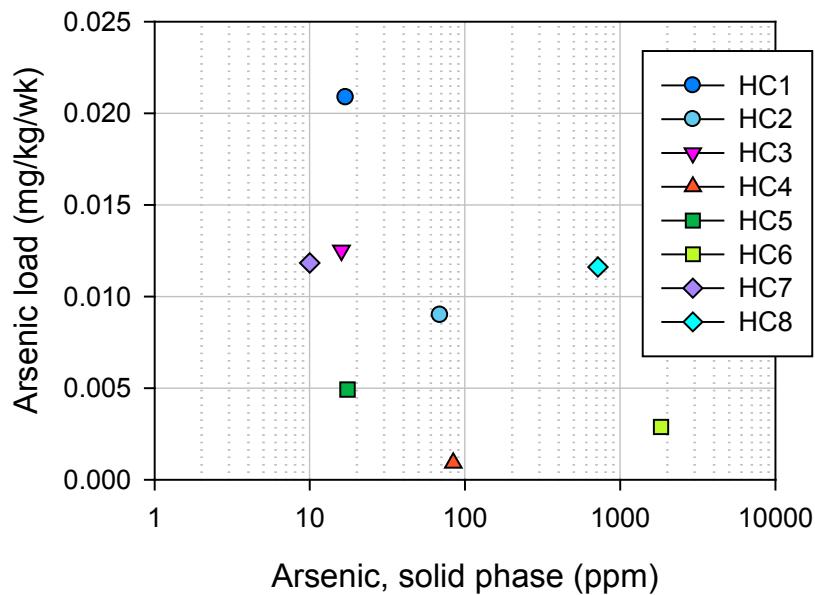
As and Mn have the highest loading rates of the selected elements, both with rates of up to 0.06 mg/kg/wk in one or more of the humidity cells in the early weeks of the tests. The initial flush of As reaches the highest rates in leachate from HC8 (up to 0.062 mg/kg/wk), although the rate decreases over time and is comparable to the rates in the other humidity cells by Week 17 (approximately 0.01 mg/kg/wk). HC1, HC2, HC3 and HC7 show gradually increasing As loading rates until approximately week 5 and have since shown a gradual decrease to more stable rates (Figure 4-12). Arsenic loading rates in HC4 have shown a consistent decrease over time from 0.0088 to 0.00068 mg/kg/wk. The two GA humidity cells (HC5 and HC6) release relatively stable and low As loading rates (<0.01 mg/kg/wk). The As loading rates do not show a correlation with the As content in the solid phase (Figure 4-13). The humidity cell with the highest average loading rate in the last four cycles of the test (HC1; 0.021 mg/kg), had relatively low As in the solid phase (17 ppm). In contrast, HC6 is composed of two high As samples (1,915 ppm and 1,750 ppm) and has a relatively low As average loading rate throughout the duration of the test (0.0030 mg/kg). It is worth noting that sample HC8, which produces the highest initial As loading rates, is the only sample in which arsenopyrite was identified on a thin section scale. Assuming that the sample splitting was representative, these trends suggest that the As association within the different humidity cell subsamples has a strong control over the respective leaching rates.



**Figure 4-11:** Relationship (a) and results (b) of model to estimate the time to onset of acidic conditions in Beaver Dam PAG samples



**Figure 4-12: D-As, D-Cu, D-Mn and D-Pb in leachate from the Beaver Dam humidity cells**



**Figure 4-13: Average As loading rate vs solid phase As content. As load is an average from weeks 17-23.**

Mn loading rates show an initial increase in the first few weeks in the majority of the humidity cell tests. The Mn loading rates are highest for leachate from HC2, HC4 and HC6. These are all the high S humidity cells for their respective lithologies. After Week 5, the Mn loading rates remain  $\leq 0.01$  mg/kg/wk for all humidity cells. Since Week 5, Mn loading rates for HC4 have been the highest. Mn loading rates from HC1 and HC7 remain low ( $< 0.002$  mg/kg/wk) over the course of the humidity cell tests. The loading rates for Cu and Pb are low ( $< 0.005$  mg/kg/wk) and do not display any obvious trends.

## 4.2 Overburden Samples

The paste pH values of the overburden samples are slightly acidic and range from pH 5.0 to pH 6.1 (Table 4-8; Appendix 4-2). One sample (LX-BDT-03) is currently acidic ( $pH < 5.5$ ). The total S content of the overburden samples is low ( $< 0.1\%$ ), with only sample LX-BDT-03 having a total S value above 0.05%. The majority of the S is present as sulphide S. This parameter is not measured but is calculated by subtracting the sulphate S, which is below the detection limit for these samples ( $< 0.01\%$ ). There was no TIC detected in the overburden samples ( $< 0.05\%$ ). The CaNP calculated from the TIC and the modified NP are both low, suggesting that the NP has been depleted from the overburden. Although the samples have a low NP, the risk for ARD is low due to the low sulphur content.

**Table 4-8:**  
**Summary of ABA results for Beaver Dam overburden samples.**

Sample ID	Paste pH	Total S	Sulphate S	Sulphide S (calc)	Total C	CaNP	Modified NP	NPR
		%	%	%	%	kg CaCO <sub>3</sub> /t	kg CaCO <sub>3</sub> /t	Mod.NP/AP
LX-BDT-01	5.6	0.02	<0.01	0.02	<0.05	<4.5	-4	-6.4
LX-BDT-02	5.7	0.02	<0.01	0.02	<0.05	<4.5	-5	-8
LX-BDT-03	5	0.08	<0.01	0.08	<0.05	<4.5	-5	-2
LX-BDT-04	6.1	0.02	<0.01	0.02	<0.05	<4.5	-4	-6.4
LX-BDT-05	5.7	0.02	<0.01	0.02	<0.05	<4.5	-4	-6.4

**Notes:** Values in grey italics are below the analytical detection limit. Values were set at the limit for calculation of NP, AP and NPR values.

Sulphate S is calculated using the HCl method.

AP (acid potential) calculated using sulphide sulphur (% non-sulphate sulphur x 31.25);

CaNP (carbonate neutralization potential) calculated using total inorganic carbon (% TIC x (100.09/12.01) x 10);

Modified NP is obtained by the modified Sobek method.

NPR = neutralization potential ratio; calculated as Modified NP / AP

The results of the total solid phase elemental analysis are included in Appendix 4-3 with a summary provided in Table 4-9. Elements that are greater than 3× their respective AUCCC values (Rudnick and Gao, 2014) and greater than 2x the detection limit in at least one of the overburden samples include As and Mn (Table 4-9). Solid phase concentrations of As in 3 of the 5 samples are enriched by a factor greater than 10 above the AUCCC. The As and Mn values are also enriched in the waste rock.

Shake flask extraction (SFE) results for the five overburden samples are presented in Appendix 4-6 and summarized in Table 4-10 as compared to the CCME water quality guidelines for the protection of aquatic life (CCME, 2018). These guidelines are used for reference only to provide an indication of parameters which are of potential concern in runoff from the excavated material. It should be noted that, as per the test method protocol, the agitation of the overburden samples in water may release higher concentrations of certain species than what would be expected in contact water drainage.

**Table 4-9:**  
**Summary of solid phase element results for Beaver Dam overburden samples.**

AUCCC	0.053	4.8	2.57	17.3	28	3.92	77.4	47	17	67
Sample ID	Ag	As	Ca	Co	Cu	Fe	Mn	Ni	Pb	Zn
	ppm	ppm	%	ppm	ppm	%	ppm	ppm	ppm	ppm
LX-BDT-01	0.2	225	0.06	13	27	3.88	490	33	13	74
LX-BDT-02	<0.2	24	0.04	5	16	2.63	245	17	9	28
LX-BDT-03	0.2	54	0.13	5	12	1.34	162	22	10	37
LX-BDT-04	<0.2	15	0.02	2	4	0.96	112	4	7	10
LX-BDT-05	0.2	105	0.05	10	15	4.51	441	19	9	64

**Notes:** Values in grey italics are below the analytical detection limit. Values were set at the limit for calculation of statistical distributions.

AUCCC = average upper continental crust concentrations (Rudnick and Gao, 2014).

Values greater than 3x the AUCCC are shaded in light grey; values greater than 10x the AUCCC are shaded in dark grey.

**Table 4-10:**  
**Summary of SFE results for Beaver Dam overburden samples.**

Parameter	Units	CCME WQG		LX-BDT-01	LX-BDT-02	LX-BDT-03	LX-BDT-04	LX-BDT-05
		Short Term	Long Term					
pH	pH	6.5-9	-	5.27	5.15	4.88	4.54	5.83
Conductivity	µS/cm	-	-	12.6	15.7	89.1	36.7	19.5
Sulphate	mg/L	-	-	2.24	1.23	10.1	1.42	2.08
<i>Leachable Metals</i>								
Ag	mg/L	-	0.00025	<0.000050	<0.000050	<0.000050	<0.000050	<0.000050
Al <sup>a</sup>	mg/L	-	0.1	<b>0.108</b>	<b>0.683</b>	<b>1.51</b>	<b>1.26</b>	<b>0.483</b>
As	mg/L	-	0.005	0.0012	0.0012	<b>0.315</b>	0.0038	<b>0.0084</b>
B	mg/L	29	1.5	<0.010	<0.010	0.015	<0.010	<0.010
Cd	mg/L	0.001	0.00009	<0.000050	<0.000050	<b>0.000217</b>	0.000075	<0.000050
Co	mg/L	-	-	0.00084	0.00042	0.00761	0.00054	0.00101
Cr	mg/L	-	0.001	<0.00050	0.00088	<b>0.00118</b>	<b>0.00126</b>	0.00064
Cu <sup>b</sup>	mg/L	-	0.002	<0.0010	<0.0010	<b>0.0085</b>	0.0014	<b>0.0033</b>
Fe	mg/L	-	0.3	<0.030	0.194	<b>0.425</b>	<b>0.490</b>	0.262
Hg	mg/L	-	0.000026	<0.000050	<0.000050	<0.000050	<0.000050	<0.000050
Mn	mg/L	-	-	0.107	0.0446	0.131	0.0681	0.106
Mo	mg/L	-	0.073	<0.00010	<0.00010	0.00023	<0.00010	<0.00010
Ni <sup>b</sup>	mg/L	-	0.025	0.00098	0.00118	<b>0.0360</b>	0.00151	0.00179
Pb <sup>b</sup>	mg/L	-	0.001	<0.00010	0.00013	0.00090	<b>0.00129</b>	0.00012
Se	mg/L	-	0.001	<0.00050	0.00068	0.00079	<b>0.00101</b>	0.00085
Tl	mg/L	-	0.0008	<0.00010	<0.00010	<0.00010	<0.00010	<0.00010
U	mg/L	0.033	0.015	0.000020	0.000087	0.000101	0.000073	0.000103
Zn	mg/L	0.037	0.007	<0.010	<0.010	<b>0.045</b>	<b>0.012</b>	<0.010

**Notes:** Values shaded in grey are above the short-term CCME guideline

Values in bold italics are above the long-term CCME guideline

<sup>a</sup>Aluminum guideline is based on pH > 6.5

<sup>b</sup>Hardness dependent guidelines are based on a hardness of 10 mg/L.

Parameters that were elevated relative to the CCME guidelines in the leachate from the samples (n=5) include Al (5 samples), As (2 samples), Cd (1 sample), Cr (2 samples), Cu (2 samples), Fe (2 samples), Pb (1 sample), Ni (1 sample), Se (1 sample) and Zn (2 samples). In addition, all samples have pH values less than 6.5 and four of the samples are acidic (pH < 5.5). Metals enriched above 10 times their respective guidelines include Al in two samples and As in one sample. Al is highly sensitive to pH and the elevated Al concentrations are likely related to the low in-situ pH of the samples. The highest As concentration in the SFE tests was measured in sample LX-BDT-03. This sample produced

As concentrations of 0.315 mg/L, while the other four overburden samples produced As concentrations ranging from 0.0012 to 0.0084 mg/L. Arsenic was identified as an element enriched in the overburden by the solid phase analysis, however, sample LX-BDT-03 does not have the highest As concentration in the solid phase. The higher As in the leachate from this sample is likely due to the low pH which would result in increased leaching from the overburden material. Given the highly weathered nature of overburden significant metal leaching potential is not expected from this unit. The cause of the elevated As concentration in LX-BDT-03 is unclear, but shows that further overburden characterization may be required.

#### 4.3 Haul Road Samples

The samples collected along the Beaver Dam haul road have paste pH values ranging from 7.4 to 8.8 and are not currently acid generating. All total S and TIC values are below their respective detection limits for these samples (Table 4-11; Appendix 4-2). The NPR values are high due to the low acid potential, indicating that the samples are not potentially acid generating. The low total S and TIC values suggest that the samples are weathered. Fresher samples could be obtained in the future to verify these results if drilling is conducted along the road alignment.

The solid phase elements are compared to the AUCCC (Rudnick and Gao, 2014; Appendix 4-3). The only elements that are above 3x the AUCCC and greater than 2x the detection limit are As (2 samples) and Mn (4 samples) (Table 4-12). Both As and Mn were also identified as elements enriched in the mine rock. As discussed above (Section 4.1.1.3), elements enriched in the solid phase may not necessarily become a metal leaching issue as there are many factors that influence leaching rates.

**Table 4-11:**  
**Summary of ABA results for Beaver Dam haul road samples.**

Sample ID	Paste pH	Total S	Sulphate S	Sulphide S (calc)	Total C	CaNP	Modified NP	NPR
		%	%	%	%	kg CaCO <sub>3</sub> /t	kg CaCO <sub>3</sub> /t	Mod.NP/AP
LX-BDR-01	7.4	<0.01	<0.01	<0.01	<0.05	<4.5	3	9.6
LX-BDR-02	8.7	<0.01	<0.01	<0.01	<0.05	<4.5	5	16.0
LX-BDR-03	8.7	<0.01	0.01	<0.01	<0.05	<4.5	4	12.8
LX-BDR-04	8.8	<0.01	0.01	<0.01	<0.05	<4.5	4	12.8
LX-BDR-05	7.7	<0.01	0.01	<0.01	<0.05	<4.5	8	25.6

**Notes:** Values in grey italics are below the analytical detection limit. Values were set at the limit for calculation of NP, AP and NPR values.

Sulphate S is calculated using the HCl method.

AP (acid potential) calculated using sulphide sulphur (% non-sulphate sulphur x 31.25);

CaNP (carbonate neutralization potential) calculated using total inorganic carbon (% TIC x (100.09/12.01) x 10);

Modified NP is obtained by the modified Sobek method.

NPR = neutralization potential ratio; calculated as Modified NP / AP

**Table 4-12:**  
**Summary of solid phase element results for Beaver Dam haul road samples.**

AUCCC	0.053	4.8	2.57	17.3	28	3.92	77.4	47	17	67
<b>Sample ID</b>	<b>Ag</b>	<b>As</b>	<b>Ca</b>	<b>Co</b>	<b>Cu</b>	<b>Fe</b>	<b>Mn</b>	<b>Ni</b>	<b>Pb</b>	<b>Zn</b>
	<i>ppm</i>	<i>ppm</i>	%	<i>ppm</i>	<i>ppm</i>	%	<i>ppm</i>	<i>ppm</i>	<i>ppm</i>	<i>ppm</i>
LX-BDR-01	<0.2	21	0.14	4	9	1.92	212	15	3	38
LX-BDR-02	<0.2	<2	0.14	13	20	2.87	392	31	5	62
LX-BDR-03	<0.2	2	0.17	7	17	2.08	302	18	7	45
LX-BDR-04	<0.2	<2	0.12	8	11	2.12	336	21	<2	39
LX-BDR-05	<0.2	18	0.78	13	28	3.63	625	38	7	66

**Notes:** Values in grey italics are below the analytical detection limit. Values were set at the limit for calculation of statistical distributions.

AUCCC = average upper continental crust concentrations (Rudnick and Gao, 2014).

Values greater than 3x the AUCCC are shaded in light grey; values greater than 10x the AUCCC are shaded in dark grey.

#### **4.4 Quality Assurance and Quality Control**

Each set of samples submitted for analyses is subjected to an internal laboratory quality assurance and quality control (QA/QC) program. This program includes duplicate samples and analytical standard analysis. Any laboratory duplicate result or standard that does not adhere to the precision specifications for the different parameters triggers a re-analysis.

## ***5. Conclusions and Recommendations***

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ATLANTIC GOLD

## **5 Conclusions and Recommendations**

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The ML/ARD characterization of the Beaver Dam samples included static test characterization of the Beaver Dam mine rock, overburden from the proposed pit area and samples from along the Beaver Dam haul road. In addition, kinetic testing was undertaken on mine rock samples. The major conclusions and recommendations of the geochemical characterization include:

- Total S in the mine rock samples varies from below the detection limit (<0.01%) up to 1.7%. The AR unit has the highest median total S (0.44%) and the GW unit has the lowest median total S (0.025%). The GA and GW units have the highest median TIC (0.11% and 0.10%, respectively) and the highest median modified NP (9.0 and 14 kg CaCO<sub>3</sub>/t, respectively), while the AR and AG units have the lowest median TIC (<0.05%) and modified NP (8.5 and 8.0 kg CaCO<sub>3</sub>/t, respectively). The mine rock samples were not currently acid generating (paste pH: 7.9 to 9.5).
- All four of the main lithologies for Beaver Dam mine rock include PAG samples, however, all samples with total S less than 0.1% are NAG. The overall proportion of PAG rock identified in the static test population is 39%; AR has the highest relative proportion of PAG samples (80%) while the GW unit has the lowest (21%).
- Parameters of potential concern identified by the total solid phase elemental analysis include As, Cu, Mn and Pb. Elevated As concentrations (>10× AUCCC) occur in all Beaver Dam lithologies. Elevated Mn values (>10× AUCCC) occur in 3 of the 4 lithologies, whereas elevated Ag and Pb (>10× AUCCC) are limited to the AG samples. Cu is not above 10x the AUCCC in any of the lithologies.
- The static test results indicate that the humidity cell samples provide a good representation of the range of geochemical characteristics within each lithologic unit. However, HC1 and HC2 have low S contents as compared to the median argillite value.
- The leachate from the kinetic tests remains circumneutral for the initial 24 weeks of data available at the time of reporting. The sulphate loading rates are lowest for HC1 (AR, median S) and HC7 (GW, median S) and highest for HC4 (AG, high S). The CMR for the majority of the humidity cells is between 1 and 2 with values up to 3 in HC1 and HC8.
- As and Mn have the highest loading rates of the parameters of potential concern identified by the solid phase analysis. Arsenic loading rates are comparable in the last weeks of the available data for the AR humidity cells (HC1 and HC2), the

median S AG humidity cell (HC3), and the GW humidity cells (HC7 and HC8). Both GA humidity cells (HC5 and HC6) have lower As loading rates and As loading rates from the high S AG humidity cell (HC4) have shown a decreasing trend over the length of the test. Mn loading rates are highest for HC4 in the later weeks of the tests. Mn loading rates from HC1 and HC7 have remained low. Mn loading rates appear to be correlated with the sulphate loading rates. The loading rates for this element are typically higher for the high S pair for each lithology.

- Based on the relationship between total S content and the NP depletion rate and in consideration of Beaver Dam site climatic conditions, it is estimated that it will take at least 17 years for the Beaver Dam PAG samples to produce ARD and around 28 years for half of all PAG samples to do so. This rate may be accelerated in ARD “hotspots” within waste rock piles where increased temperatures may be attained in response to heat produced by sulphide oxidation.
- The overburden samples have low sulphur content; however, they lack any significant NP. Elements enriched in the overburden include As and Mn. The SFE tests indicate the potential for elevated Al and As in runoff, among other elements. However, significant metal leaching is not expected from this unit due to the high degree of weathering. Further characterization of the overburden material is recommended.
- Samples collected along the haul road have low total S and low NP and do not currently indicate a potential for ARD. It is assumed that these samples are surface-weathered and therefore, it is recommended that fresh, unweathered samples be collected to verify the results if drilling is conducted along the haul road. Elements of potential concern identified by the solid phase analysis include As and Mn.

## ***6. Closure***

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ATLANTIC GOLD

## **6    Closure**

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This report has been prepared for AMNS for the ML/ARD assessment of the Beaver Dam Gold project. Please contact the undersigned should you require any additional information or clarification on the contents of this report.

Sincerely,

**LORAX ENVIRONMENTAL SERVICES LTD.**

**Prepared by:**

*Original Signed By:*

**Jennifer Owen, M.Sc., GIT**  
Environmental Scientist

**Prepared by:**

*Original Signed and Sealed By:*

**Alison Shaw, Ph.D., P.Geo**  
Senior Environmental Geoscientist

**Reviewed by:**

*Original Signed By:*

**Timo Kirchner, M.Sc., P.Geo**  
Environmental Geoscientist

**Reviewed by:**

*Original Signed By:*

**Bruce Mattson, M.Sc., P.Geo**  
Senior Environmental Geoscientist

## ***References***

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## References

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- Canadian Council of Ministers of the Environment (CCME) (2018). Canadian Environmental Quality Guidelines, water quality guidelines for the protection of aquatic life, <http://ceqg-rcqe.ccme.ca/en/index.html>. Accessed November 2018.
- Dockrey, J. & Mattson, B (2016). Effects of pH on the Arrhenius Paradigm. In Proceedings of the International Mine Water Association (IMWA) Annual Conference, Leipzig, Germany, July 11-15, 2016.
- FSSI Consultants (2015). Technical Report of the Beaver Dam Gold Project, Halifax County, Nova Scotia. Prepared by Neil Schofield of FSS International Consultants (Australia) Pty Ltd., prepared for Atlantic Gold Corporation, dated March 2, 2015.
- Government of Canada (2018). Canadian Climate Normals: 1981-2010. [climate.weather.gc.ca/climate\\_normals/index\\_e.html](http://climate.weather.gc.ca/climate_normals/index_e.html). Accessed November 6, 2018.
- Jambor, J., Dutrizac, J., Raudsepp, M., and Groat, L. (2003). Effect of Peroxide on Neutralization Potential Values of Siderite and Other Carbonate Minerals. *Journal of Environmental Quality* 32, 2373-2378.
- Lapakko, K. A. (2003). Developments in humidity cell tests and their application, in: Jambor, J. L.; Blowes, D. W.; Ritchie, A. I. M., (Eds.), *Environmental aspects of mine wastes*, Minerals Association of Canada 147-164.
- Lorax Environmental Services Ltd. (Lorax) (2018). Beaver Dam Project - Geochemical Source Term Predictions for Waste Rock, Low-Grade Ore, Tailings and Overburden. Draft Technical Report submitted to AMNS on November 8, 2018.
- Mattson, B. (2005). Evaluating Depletion of Carbonate Neutralization Potential from Laboratory Kinetic Tests. *Securing the Future 2005, International Conference on Mining and the Environment and Metals and Energy Recovery. June 27 - July 1, 2005, Skellefteå, Sweden.*
- Price, W.A. (1997). Draft Guidelines and Recommended Methods for the Prediction of Metal Leaching and Acid Rock Drainage at Mine-sites in British Columbia. Reclamation Section, Energy and Minerals Division, British Columbia Ministry of Employment and Investment.
- Price, W.A. (2009). MEND Report 1.20.1 – Prediction Manual for Drainage Chemistry from Sulphidic Geologic Materials, dated December 2009.

- Rudnick, R.L. and S. Gao (2014). Composition of the Continental Crust. In: Holland, H. and Turekian, K. (eds). Treatise on Geochemistry 2nd Edition, Vol. 4, pp. 1-51. Oxford, UK, Elsevier Ltd.
- Sapsford, D.J., Bowell, R.J., Dey, M., and Williams, K.P. (2009). Humidity cell tests for the prediction of acid rock drainage. Minerals Engineering 22, 25-36.

## **Appendices**

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***Appendix 3-1:  
Sample Locations and  
Drill Core Details***

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**Appendix 3-1 Sample Locations and Drill Core Details**

Geologic Unit	Hole ID	Interval (m)		Sample ID	Geologic Unit	Hole ID	Interval (m)		Sample ID
		From	To				From	To	
Argillite (AR)	BD14-178	7	7.9	LX17-12	Greywacke (GW)	BD15-GT08	91	10	LX17-06
	BD15-GT05	57	58	LX17-22		BD15-GT08	23	24	LX17-08
	BD14-188	22	23	LX17-24		BD14-178	49	50	LX17-16
	BD15-190	40	41	BD190-040		BD15-GT05	25	26	LX17-20
	BD05-005	54	55	BD005-054		BD15-GT05	41	42	LX17-21
	BD06-110	50	51	BD110-050		BD14-156	75	76	BD156-075
	BD15-190	43	44	BD190-043		BD14-160	39	40	BD160-039
	BD14-157	31	32	BD157-031		BD14-179	45	46	BD179-045
	BD06-085	46	47	BD085-046		BD14-179	149	150	BD179-149
	BD05-011	123	124	BD011-123		BD05-006	12	13	BD006-012
Argillite-Greywacke (AG)	BD15-GT02	15	16	LX17-02		BD14-169	44	45	BD169-044
	BD15-GT02	26	27	LX17-03		BD14-169	88	89	BD169-088
	BD15-GT02	31	32	LX17-04		BD05-011	15	16	BD011-015
	BD15-GT08	37	38	LX17-09		BD14-183	70	71	BD183-070
	BD14-172	140	141	LX17-10		BD14-183	148	149	BD183-148
	BD14-172	170	171	LX17-11		BD06-049	140	141	BD049-140
	BD14-178	15	16	LX17-13		BD14-179	43	44	BD179-043
	BD14-173	22	23	LX17-28		BD05-006	13	14	BD006-013
	BD14-160	140	141	BD160-140		BD14-183	132	133	BD183-132
	BD14-179	170	171	BD179-170		BD14-186	149	150	BD186-149
Greywacke-Agillite (GA)	BD14-169	136	137	BD169-136		BD14-188	38	39	LX17-25
	BD15-GT02	10	11	LX17-01		BD14-188	51	52	LX17-26
	BD15-GT02	46	47	LX17-05		BD14-173	12	13	LX17-27
	BD15-GT08	14	15	LX17-07		BD14-173	53	54	LX17-30
	BD14-178	30	31	LX17-14	Quartz Vein (QTZV)	BD14-157	70	71	BD157-070
	BD14-178	58	59	LX17-15		BD14-156	108	109	BD156-108
	BD14-178	161	162	LX17-17		BD14-160	89	90	BD160-089
	BD14-178	147	148	LX17-18					
	BD15-GT05	15	16	LX17-19					
	BD14-188	10	11	LX17-23					
	BD14-173	37	38	LX17-29					
	BD14-156	9	10	BD156-009					
	BD14-187A	10	11	BD187A-010					
	BD14-157	33	34	BD157-033					
	BD14-169	187	188	BD169-187					

## ***Appendix 4-1: XRD Results***

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ATLANTIC GOLD

**QUANTITATIVE PHASE ANALYSIS OF 8 POWDER SAMPLES USING THE RIETVELD METHOD AND X-RAY POWDER DIFFRACTION DATA.**

**Project: 1821 Beaver Dam  
PO:43079**

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**Rik Vos - Carolyn Jones  
SGS Canada Inc.  
3260 Production Way  
Burnaby, BC V5A 4W4**

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**Mati Raudsepp, Ph.D.  
Elisabetta Pani, Ph.D.  
Edith Czech, M.Sc.  
Jenny Lai, B.Sc.  
Lan Kato, B.A.**

**Dept. of Earth & Ocean Sciences  
The University of British Columbia  
6339 Stores Road  
Vancouver, BC V6T 1Z4**

**July 10, 2018**

## **EXPERIMENTAL METHOD**

The samples of **Project 1821 Beaver Dam** were reduced to the optimum grain-size range for quantitative X-ray analysis ( $<10 \mu\text{m}$ ) by grinding under ethanol in a vibratory McCrone Micronizing Mill for 10 minutes. Continuous-scan X-ray powder-diffraction data were collected over a range  $3\text{-}80^\circ 2\theta$  with CoK $\alpha$  radiation on a Bruker D8 Advance Bragg-Brentano diffractometer equipped with an Fe monochromator foil, 0.6 mm ( $0.3^\circ$ ) divergence slit, incident- and diffracted-beam Soller slits and a LynxEye-XE detector. The long fine-focus Co X-ray tube was operated at 35 kV and 40 mA, using a take-off angle of  $6^\circ$ .

## **RESULTS**

The X-ray diffractograms were analyzed using the International Centre for Diffraction Database PDF-4 and Search-Match software by Bruker. X-ray powder-diffraction data of the samples were refined with Rietveld program Topas 4.2 (Bruker AXS). The results of quantitative phase analysis by Rietveld refinements are given in Table 1 (separate file, ***SGS Canada Results July 10 2018 - Project 1821 Beaver Dam - PO43079 - 8 samples.xls***). These amounts represent the relative amounts of crystalline phases normalized to 100%. The Rietveld refinement plots are shown in Figures 1 - 8. The ideal formulae of the mineral phases are shown in Table 2.

Table 2.

<b>Mineral</b>	<b>Ideal Formula</b>
Biotite	K(Mg,Fe)₃(AlSi₃O₁₀)(OH)₂
Calcite	CaCO₃
Chamosite	(Fe,Al,Mg)₆(Si,Al)₄O₁₀(OH)₈
Illite/Muscovite 2M1	K <sub>0.65</sub> Al <sub>2.0</sub> Al <sub>0.65</sub> Si <sub>3.35</sub> O <sub>10</sub> (OH) <sub>2</sub> /KAl <sub>2</sub> (AlSi <sub>3</sub> O <sub>10</sub> )(OH) <sub>2</sub>
Ilmenite ?	Fe <sup>2+</sup> TiO <sub>3</sub>
Plagioclase	NaAlSi <sub>3</sub> O <sub>8</sub> – CaAl <sub>2</sub> Si <sub>2</sub> O <sub>8</sub>
Pyrite	FeS <sub>2</sub>
Quartz	SiO <sub>2</sub>

1SGS\_HC1.raw

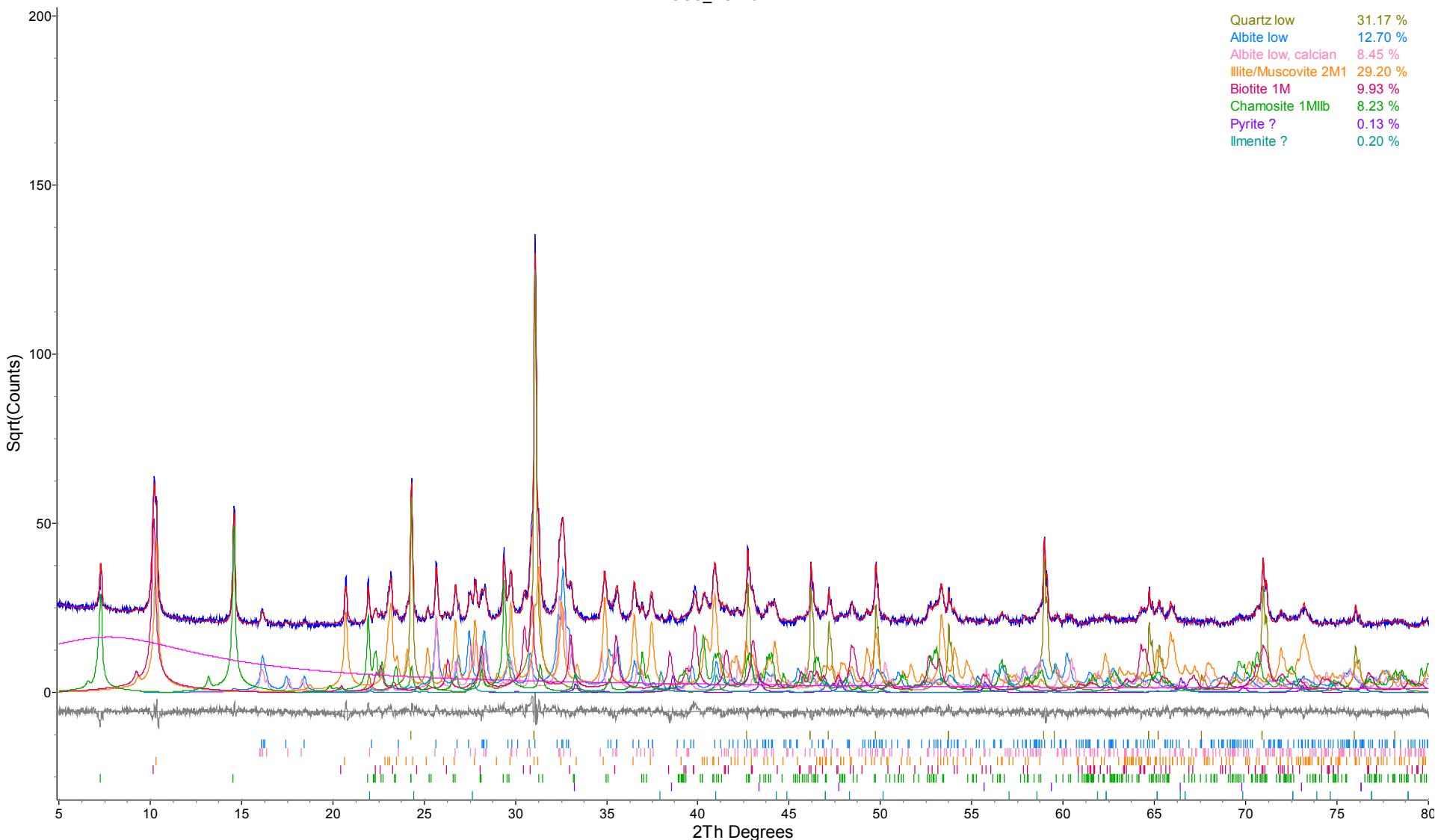


Figure 1. Rietveld refinement plot of sample **SGS CANADA HC 1** (blue line - observed intensity at each step; red line - calculated pattern; solid grey line below – difference between observed and calculated intensities; vertical bars, positions of all Bragg reflections). Coloured lines are individual diffraction patterns of all phases.

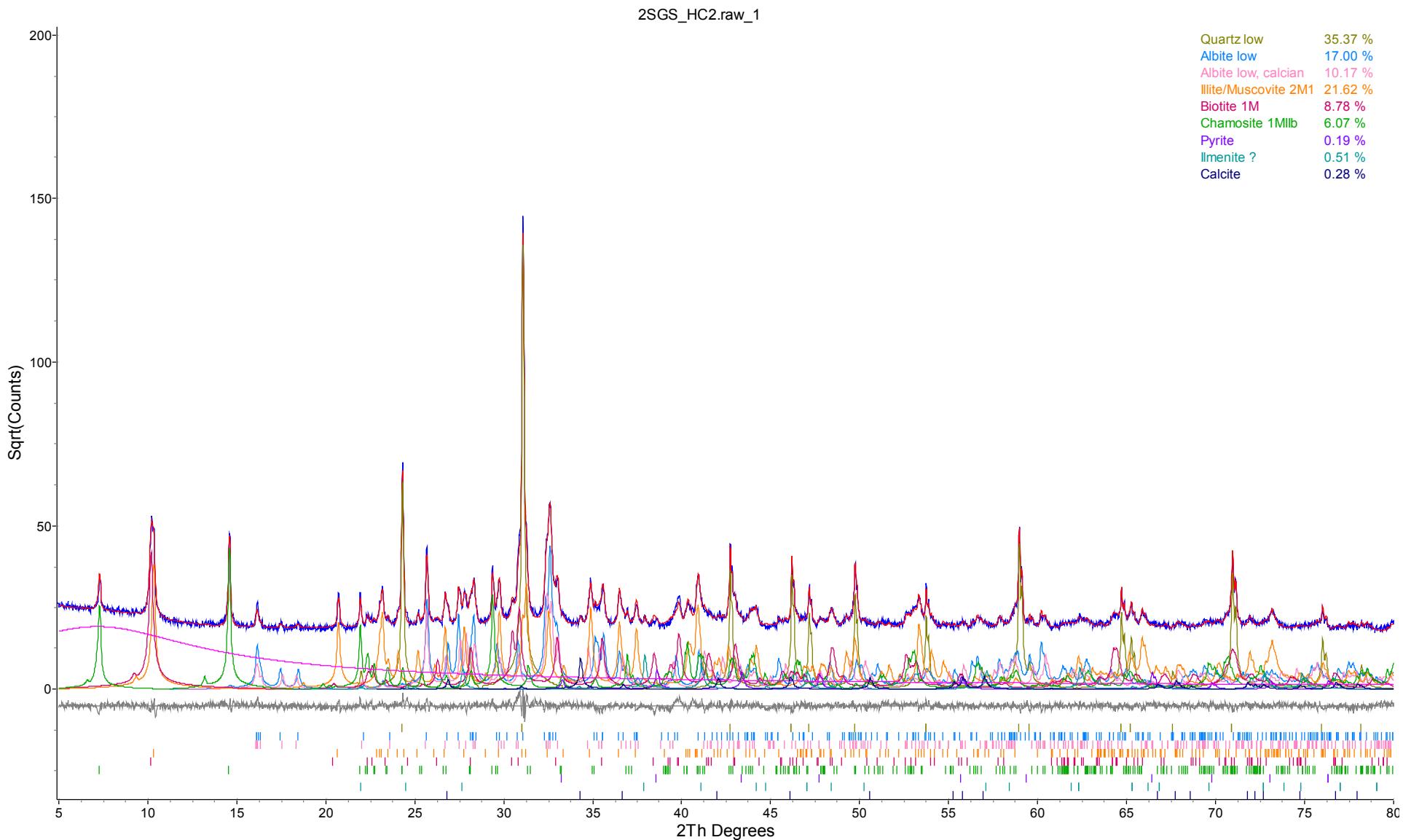


Figure 2. Rietveld refinement plot of sample **SGS CANADA HC 2** (blue line - observed intensity at each step; red line - calculated pattern; solid grey line below - difference between observed and calculated intensities; vertical bars - positions of all Bragg reflections). Coloured lines are individual diffraction patterns of all phases.

3SGS\_HC3.raw

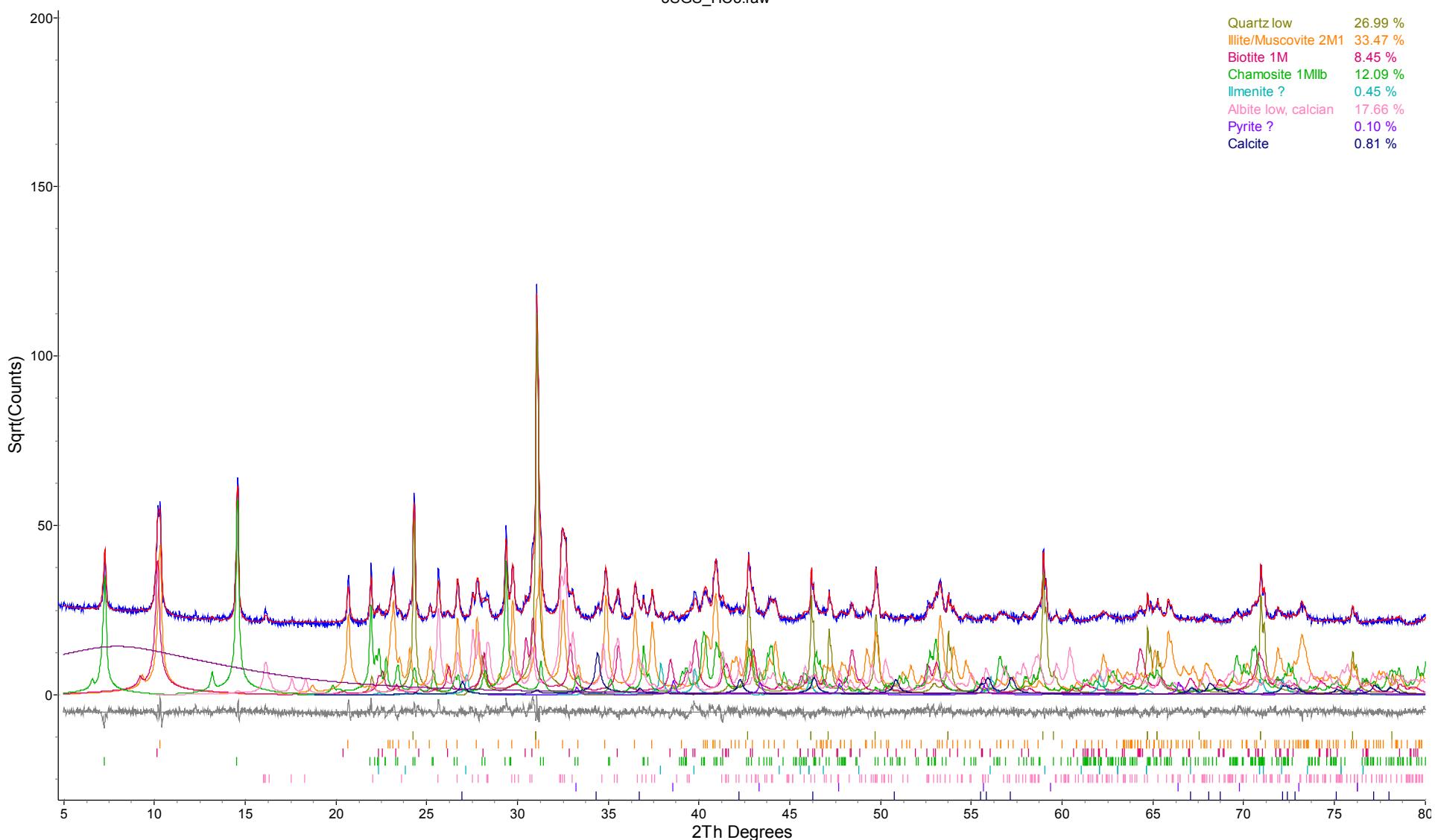


Figure 3. Rietveld refinement plot of sample **SGS CANADA HC 3** (blue line - observed intensity at each step; red line - calculated pattern; solid grey line below - difference between observed and calculated intensities; vertical bars - positions of all Bragg reflections). Coloured lines are individual diffraction patterns of all phases.

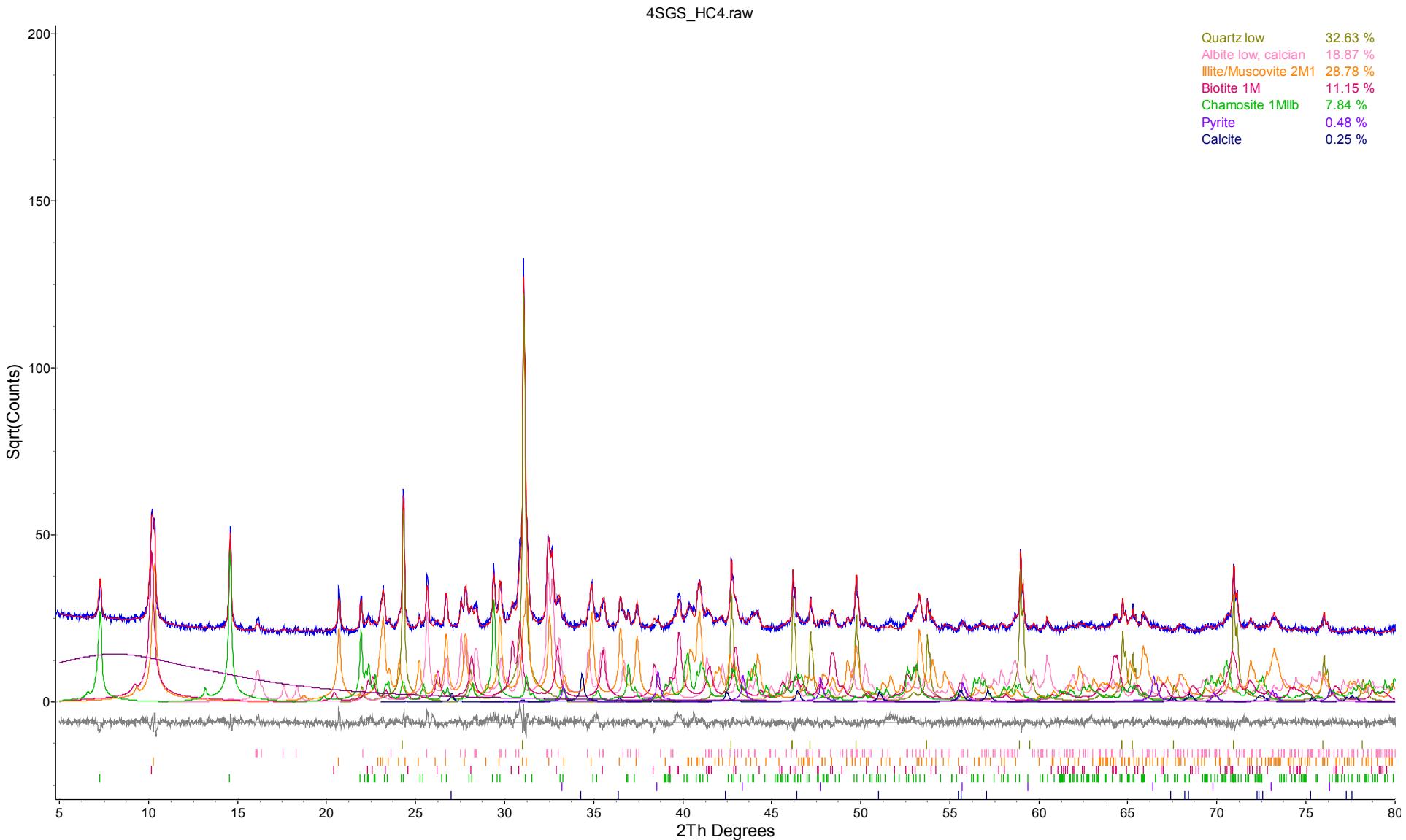


Figure 4. Rietveld refinement plot of sample **SGS CANADA HC 4** (blue line - observed intensity at each step; red line - calculated pattern; solid grey line below - difference between observed and calculated intensities; vertical bars - positions of all Bragg reflections). Coloured lines are individual diffraction patterns of all phases.

5SGS\_HC5.raw

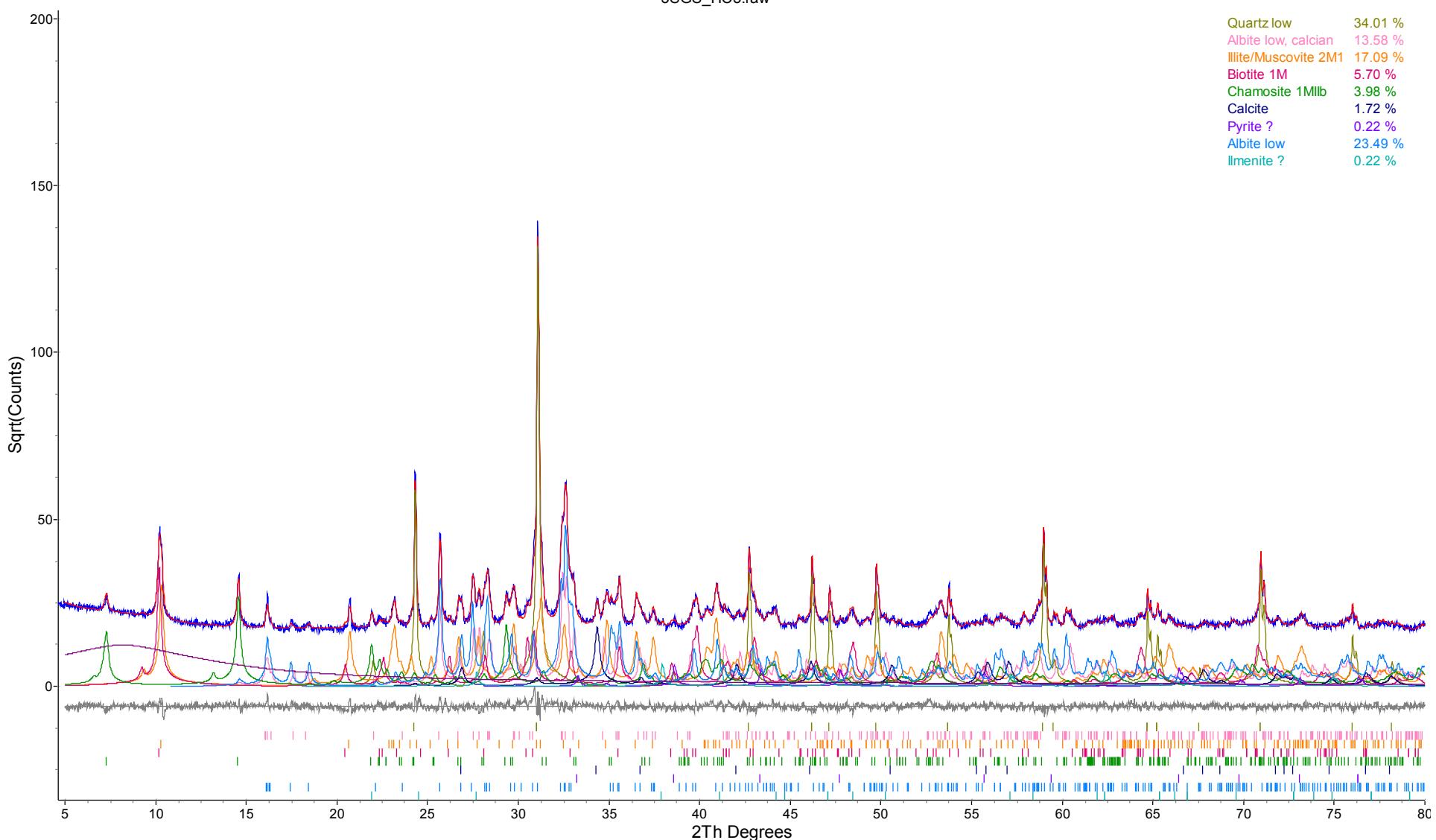


Figure 5. Rietveld refinement plot of sample **SGS CANADA HC 5** (blue line - observed intensity at each step; red line - calculated pattern; solid grey line below - difference between observed and calculated intensities; vertical bars - positions of all Bragg reflections). Coloured lines are individual diffraction patterns of all phases.

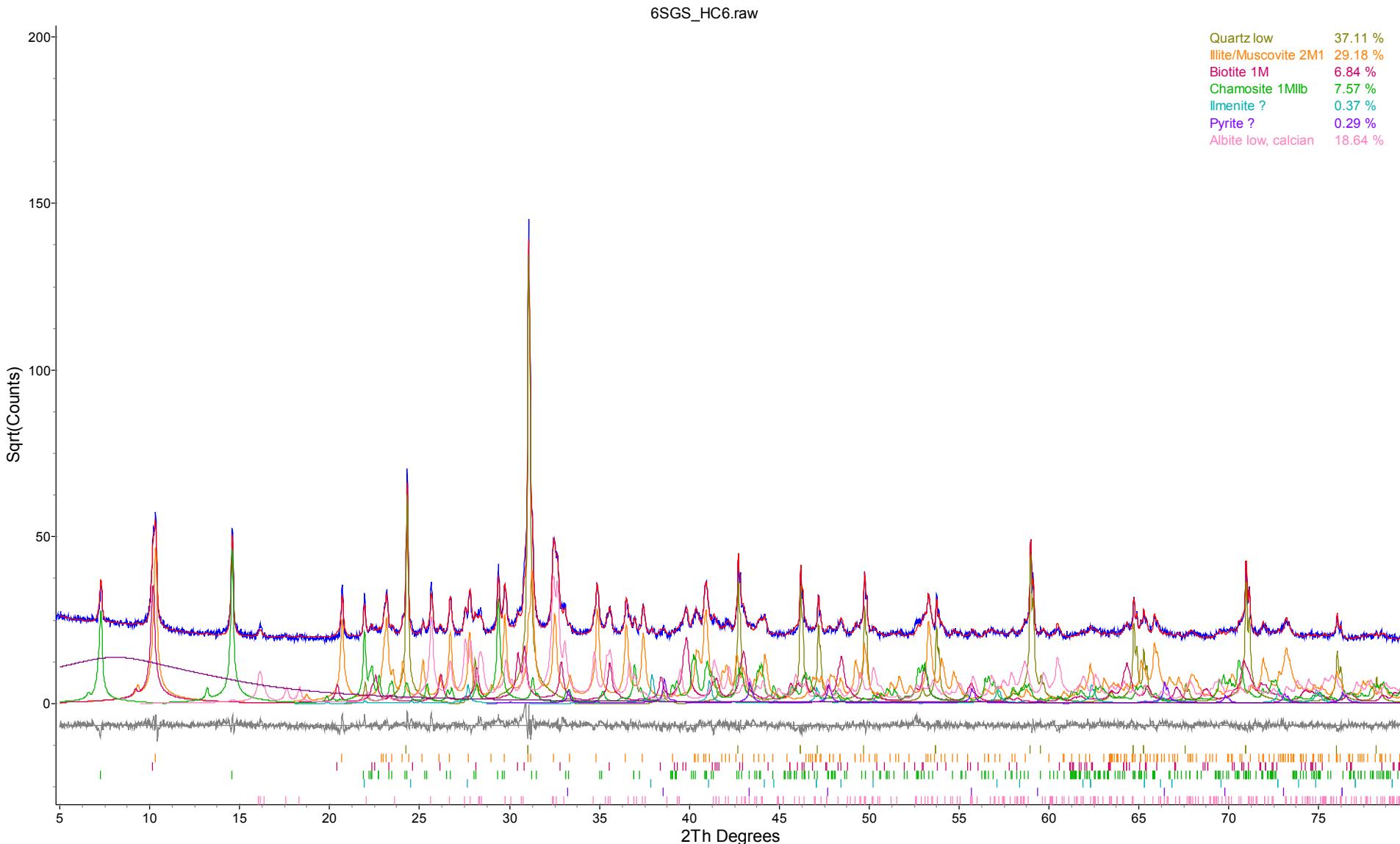


Figure 6. Rietveld refinement plot of sample **SGS CANADA HC 6** (blue line - observed intensity at each step; red line - calculated pattern; solid grey line below - difference between observed and calculated intensities; vertical bars - positions of all Bragg reflections). Coloured lines are individual diffraction patterns of all phases.

7SGS\_HC7.raw

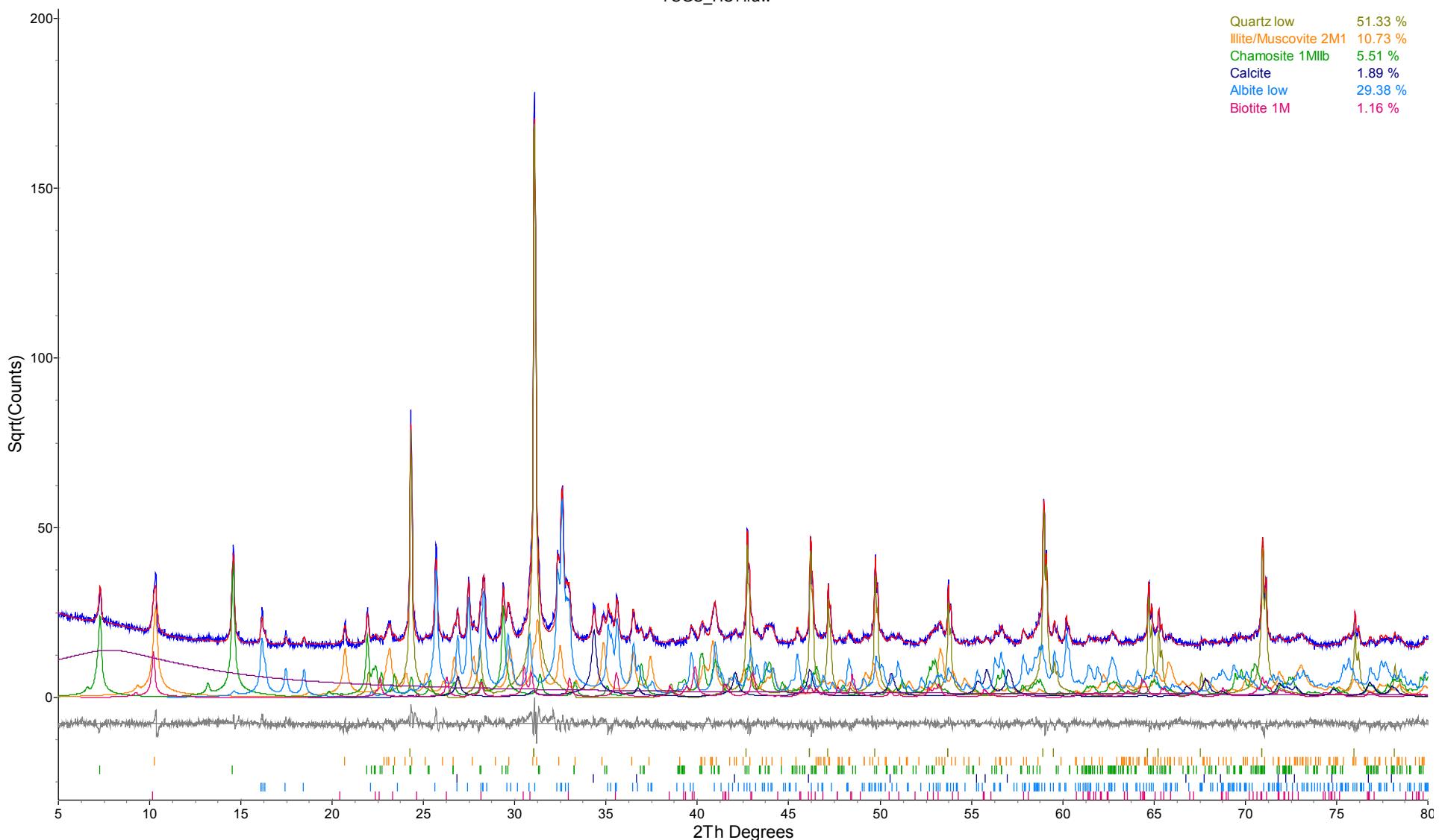


Figure 7. Rietveld refinement plot of sample **SGS CANADA HC 7** (blue line - observed intensity at each step; red line - calculated pattern; solid grey line below - difference between observed and calculated intensities; vertical bars - positions of all Bragg reflections). Coloured lines are individual diffraction patterns of all phases.

8SGS\_HC8.raw

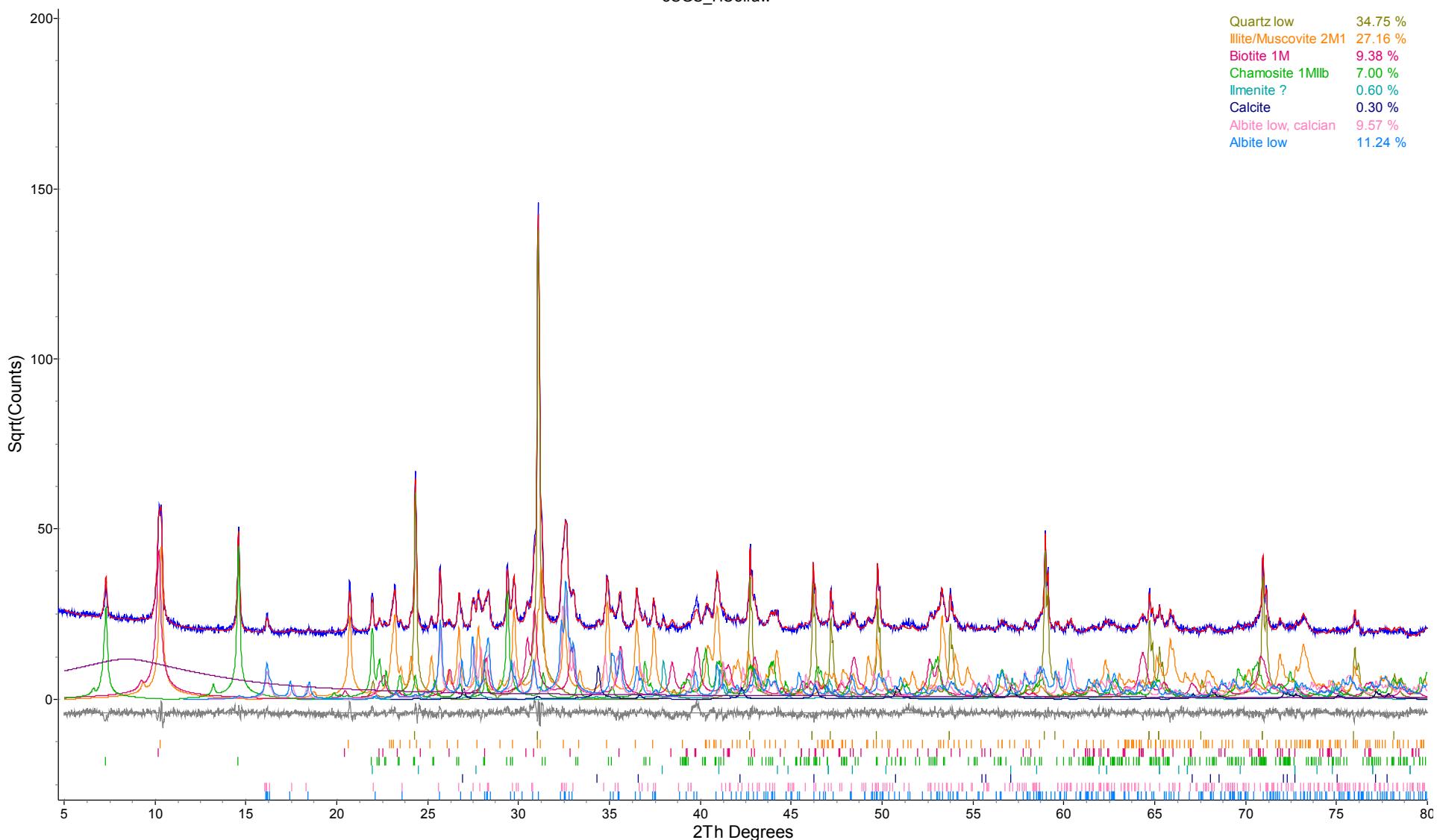


Figure 8. Rietveld refinement plot of sample **SGS CANADA HC 8** (blue line - observed intensity at each step; red line - calculated pattern; solid grey line below - difference between observed and calculated intensities; vertical bars - positions of all Bragg reflections). Coloured lines are individual diffraction patterns of all phases.

## ***Appendix 4-2: Acid-Base Accounting Results***

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ATLANTIC GOLD





**Appendix 4-2, Table 2: Acid-Base Accounting Results - Kinetic Test Samples**

Humidity Cell ID	Description	Sample ID	Hole ID	Interval		Paste pH	Total S	Sulphate S (CO <sub>3</sub> )	Sulphate S (HCl)	Sulphide S (calc)	Total C	CO <sub>2</sub>	CO <sub>3</sub>	CaNP	Modified NP	NPR	NPR
				From	To		%	%	%	%	%	%	%	kg CaCO <sub>3</sub> /t	kg CaCO <sub>3</sub> /t	CaNP/AP	Mod.NP/AP
HC1	Median-S AR	LX17-12	BD14-178	7	7.9	8.9	0.02	<0.01	0.02	0.02	<0.05	<0.2	<0.2	<4.5	6	<7.3	9.6
HC2	High-S AR	LX17-22	BD15-GT05	57	58	9.0	0.32	<0.01	0.01	0.32	<0.05	<0.2	<0.2	<4.5	6	<0.5	0.60
		LX17-24	BD14-188	22	23	8.4	0.18	<0.01	0.02	0.18	<0.05	<0.2	<0.2	<4.5	9	<0.8	1.6
Average						8.7	0.25	<0.01	0.02	0.25	<0.05	<0.2	<0.2	<4.5	8	<0.6	1.1
HC3	Median-S AG	LX17-02	BD15-GT02	15	16	8.2	0.32	0.03	0.02	0.29	<0.05	<0.2	<0.2	<4.5	7	<0.5	0.77
		LX17-10	BD14-172	140	141	9.1	0.07	<0.01	0.02	0.07	0.17	0.6	0.8	14	23	6.2	11
Average						8.7	0.20	0.02	0.02	0.18	0.11	0.4	0.5	9.1	15	3.4	5.6
HC4	High-S AG	LX17-03	BD15-GT02	26	27	8.5	0.4	0.02	0.01	0.38	<0.05	<0.2	<0.2	<4.5	8	<0.4	0.67
		LX17-04	BD15-GT02	31	32	7.9	1.1	0.01	0.02	1.1	<0.05	0.2	0.2	4.5	8	0.1	0.23
Average						8.2	0.77	0.02	0.02	0.76	<0.05	0.2	0.2	4.5	8	0.3	0.45
HC5	Median-S GA	LX17-01	BD15-GT02	10	11	7.9	0.36	0.03	0.01	0.33	0.06	0.2	0.3	4.5	9	0.4	0.87
		LX17-23	BD14-188	10	11	9.0	0.05	<0.01	0.02	0.05	0.3	1.1	1.5	25	34	16	22
Average						8.5	0.21	0.02	0.02	0.19	0.2	0.7	0.9	15	22	8.2	11.3
HC6	High-S GA	LX17-17	BD14-178	161	162	8.1	0.48	<0.01	0.01	0.48	<0.05	<0.2	<0.2	<4.5	6	<0.3	0.40
		LX17-19	BD15-GT05	15	16	9.0	0.30	<0.01	0.02	0.30	<0.05	<0.2	<0.2	<4.5	5	<0.5	0.53
Average						8.6	0.39	<0.01	0.02	0.39	<0.05	<0.2	<0.2	<4.5	6	0.4	0.47
HC7	Median-S GW	LX17-16	BD14-178	49	50	9.1	0.03	<0.01	<0.01	0.03	0.2	0.7	1.0	16	23	17	25
HC8	High-S GW	LX17-20	BD15-GT05	25	26	9.0	0.29	0.01	0.03	0.28	<0.05	<0.2	<0.2	<4.5	7	<0.5	0.80
		LX17-21	BD15-GT05	41	42	9.5	0.42	<0.01	0.03	0.42	0.05	0.2	0.2	4.5	9	0.3	0.69
Average						9.3	0.36	0.01	0.03	0.35	0.05	0.2	0.2	4.5	8	0.4	0.74

**Notes:**

For humidity cells made up of two samples, a 1:1 mixture is assumed.

A hyphen indicates the parameter was not analyzed.

Values in grey italics are below the analytical detection limit. Values were set at the limit for calculation of NP, AP and NPR values.

Sulphate S is calculated using the sodium carbonate leach method (CO<sub>3</sub>) and the HCl method (HCl).

AP (acid potential) calculated using sulfide sulfur (% non-sulfate sulfur x 31.25)

CaNP (carbonate neutralization potential) calculated using total inorganic carbon (% TIC x (100.09/12.01) x 10)

Modified NP is obtained by modified Sobek method.

NPR = neutralization potential ratio; calculated as CaNP / AP and as Modified NP / AP.

**Appendix 4-2, Table 3: Acid-Base Accounting Results - Overburden**

Sample ID	Paste pH	Total S	Sulphate S (CO <sub>3</sub> )	Sulphate S (HCl)	Sulphide S (calc)	Total C	CO <sub>2</sub>	CaNP	Modified NP	NPR	NPR
		%	%	%	%	%	%	kg CaCO <sub>3</sub> /t	kg CaCO <sub>3</sub> /t	CaNP/AP	Mod.NP/AP
LX-BDT-01	5.6	0.02	<0.01	<0.01	0.02	<0.05	<0.2	<4.5	-4	7.28	-6.4
LX-BDT-02	5.7	0.02	<0.01	<0.01	0.02	<0.05	<0.2	<4.5	-5	7.28	-8
LX-BDT-03	5	0.08	<0.01	<0.01	0.08	<0.05	<0.2	<4.5	-5	1.82	-2
LX-BDT-04	6.1	0.02	<0.01	<0.01	0.02	<0.05	<0.2	<4.5	-4	7.28	-6.4
LX-BDT-05	5.7	0.02	<0.01	<0.01	0.02	<0.05	<0.2	<4.5	-4	7.28	-6.4

**Notes:** Values in grey italics are below the analytical detection limit. Values were set at the limit for calculation of NP, AP and NPR values.

Sulphate S is calculated using the sodium carbonate leach method (CO<sub>3</sub>) and the HCl method (HCl).

AP (acid potential) calculated using sulfide sulfur (% non-sulfate sulfur x 31.25)

CaNP (carbonate neutralization potential) calculated using total inorganic carbon (% TIC x (100.09/12.01) x 10)

Modified NP is obtained by modified Sobek method.

NPR = neutralization potential ratio; calculated as CaNP / AP and as Modified NP / AP.

**Appendix 4-1, Table 4: Acid-Base Accounting Results - Haul Road**

Sample ID	Paste pH	Total S	Sulphate S (CO <sub>3</sub> )	Sulphate S (HCl)	Sulphide S (calc)	Total C	CO <sub>2</sub>	CaNP	Modified NP	NPR	NPR
		%	%	%	%	%	%	kg CaCO <sub>3</sub> /t	kg CaCO <sub>3</sub> /t	CaNP/AP	Mod.NP/AP
LX-BDR-01	7.4	<0.01	<0.01	<0.01	<0.01	<0.05	<0.2	<4.5	3	14.6	9.6
LX-BDR-02	8.7	<0.01	<0.01	<0.01	<0.01	<0.05	<0.2	<4.5	5	14.6	16.0
LX-BDR-03	8.7	<0.01	<0.01	0.01	<0.01	<0.05	<0.2	<4.5	4	14.6	12.8
LX-BDR-04	8.8	<0.01	<0.01	0.01	<0.01	<0.05	<0.2	<4.5	4	14.6	12.8
LX-BDR-05	7.7	<0.01	<0.01	0.01	<0.01	<0.05	<0.2	<4.5	8	14.6	25.6

**Notes:** Values in grey italics are below the analytical detection limit. Values were set at the limit for calculation of NP, AP and NPR values.

Sulphate S is calculated using the sodium carbonate leach method (CO<sub>3</sub>) and the HCl method (HCl).

AP (acid potential) calculated using sulfide sulfur (% non-sulfate sulfur x 31.25)

CaNP (carbonate neutralization potential) calculated using total inorganic carbon (% TIC x (100.09/12.01) x 10)

Modified NP is obtained by modified Sobek method.

NPR = neutralization potential ratio; calculated as CaNP / AP and as Modified NP / AP.

## ***Appendix 4-3: Solid Phase Elemental Analysis Results***

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ATLANTIC GOLD





**Appendix 4-3, Table 2: Solid Phase Elemental Analysis Results - Humidity Cell Samples**

Humidity Cell ID	Description	Sample ID	Ag	Al	As	B	Ba	Be	Bi	Ca	Cd	Co	Cr	Cu	Fe	Hg	Hg	K	La
			ppm	%	ppm	ppm	ppm	ppm	ppm	%	ppm	ppm	ppm	ppm	%	ppm	ppm	%	ppm
HC1	Median-S AR	LX17-12	<0.2	2.74	17	<10	90	0.5	<2	0.20	<0.5	20	41	28	4.57	<1	<0.005	1.08	30
HC2	High-S AR	LX17-22	<0.2	1.24	116	<10	90	<0.5	<2	0.15	<0.5	9	36	18	2.48	<1	<0.005	0.64	10
		LX17-24	<0.2	2.94	22	<10	110	0.7	<2	0.36	<0.5	19	41	44	4.85	<1	<0.005	0.99	30
Average			<0.2	2.09	69	<10	100	0.6	<2	0.26	<0.5	14	39	31	3.67	<1	<0.005	0.82	20
HC3	Median-S AG	LX17-02	<0.2	3.22	17	<10	80	0.5	<2	0.20	<0.5	24	37	42	6.22	<1	<0.005	0.83	40
		LX17-10	<0.2	2.64	15	<10	120	<0.5	<2	0.85	<0.5	19	48	4	4.45	1	<0.005	0.85	20
	Average		<0.2	2.93	16	<10	100	0.5	<2	0.53	<0.5	22	43	23	5.34	1	<0.005	0.84	30
HC4	High-S AG	LX17-03	<0.2	2.72	157	<10	110	<0.5	<2	0.26	<0.5	19	44	55	5.3	<1	<0.005	0.98	30
		LX17-04	6.6	2.72	11	<10	110	0.6	26	0.25	0.5	29	43	155	6.16	<1	<0.005	1.11	30
	Average		3.4	2.72	84	<10	110	0.6	14	0.26	0.5	24	44	105	5.73	<1	<0.005	1.05	30
HC5	Median-S GA	LX17-01	0.5	1.5	10	<10	60	0.5	<2	0.40	<0.5	13	38	47	3.06	<1	<0.005	0.43	30
		LX17-23	<0.2	1.99	25	<10	130	0.6	<2	1.33	<0.5	13	41	25	3.28	<1	<0.005	0.98	30
	Average		0.4	1.75	18	<10	95	0.6	<2	0.87	<0.5	13	40	36	3.17	<1	<0.005	0.71	30
HC6	High-S GA	LX17-17	<0.2	1.97	1915	<10	120	0.6	<2	0.26	<0.5	14	38	51	3.63	<1	<0.005	0.78	20
		LX17-19	<0.2	2.71	1750	<10	90	<0.5	<2	0.12	<0.5	25	39	21	5.12	<1	<0.005	0.86	30
	Average		<0.2	2.34	1833	<10	105	0.6	<2	0.19	<0.5	20	39	36	4.38	<1	<0.005	0.82	25
HC7	Median-S GW	LX17-16	<0.2	1.13	10	<10	40	0.7	<2	0.93	<0.5	7	31	18	2.12	<1	<0.005	0.21	30
HC8	High-S GW	LX17-20	<0.2	2.68	231	<10	80	<0.5	<2	0.16	<0.5	25	37	35	5.23	<1	<0.005	0.85	20
		LX17-21	0.3	1.61	1205	<10	130	<0.5	<2	0.28	<0.5	16	39	24	3.24	<1	<0.005	0.92	20
Average			0.3	2.15	718	<10	105	<0.5	<2	0.22	<0.5	21	38	30	4.24	<1	<0.005	0.89	20

**Notes:**

For humidity cells made up of two samples, a 1:1 mixture is assumed

Values in grey italics are below the analytical detection limit.

**Appendix 4-3, Table 2: Solid Phase Elemental Analysis Results - Humidity Cell Samples**

Humidity Cell ID	Description	Sample ID	Mg	Mn	Mo	Na	Ni	P	Pb	S	Sb	Sc	Sr	Th	Ti	Tl	U	V	W	Zn
			%	ppm	ppm	%	ppm	ppm	ppm	%	ppm	ppm	ppm	ppm	%	ppm	ppm	ppm	ppm	
HC1	Median-S AR	LX17-12	1.43	534	<1	0.03	44	610	11	0.02	<2	6	10	<20	0.17	<10	<10	44	<10	89
HC2	High-S AR	LX17-22	0.75	312	<1	0.05	15	210	20	0.35	<2	6	5	<20	0.13	<10	<10	40	<10	57
		LX17-24	1.63	621	<1	0.03	42	650	6	0.19	<2	5	16	<20	0.15	<10	<10	45	<10	120
Average			1.19	467	<1	0.04	29	430	13	0.27	<2	6	11	<20	0.14	<10	<10	43	<10	89
HC3	Median-S AG	LX17-02	1.65	634	1	0.02	43	520	10	0.33	<2	4	4	<20	0.13	<10	<10	42	<10	123
		LX17-10	1.52	774	<1	0.05	43	640	7	0.06	<2	7	20	<20	0.16	<10	<10	54	<10	103
Average			1.59	704	1	0.04	43	580	9	0.20	<2	6	12	<20	0.15	<10	<10	48	<10	113
HC4	High-S AG	LX17-03	1.33	601	<1	0.04	37	490	11	0.43	<2	7	9	<20	0.15	<10	<10	54	<10	102
		LX17-04	1.51	675	1	0.03	59	360	179	1.26	<2	7	7	<20	0.15	<10	<10	50	<10	143
Average			1.42	638	1	0.04	48	425	95	0.85	<2	7	8	<20	0.15	<10	<10	52	<10	123
HC5	Median-S GA	LX17-01	0.8	447	15	0.06	27	560	54	0.41	<2	6	8	<20	0.09	<10	<10	40	<10	63
		LX17-23	1.05	512	<1	0.05	30	650	7	0.05	<2	7	20	<20	0.17	<10	<10	51	<10	76
Average			0.93	480	8	0.06	29	605	31	0.23	<2	7	14	<20	0.13	<10	<10	46	<10	70
HC6	High-S GA	LX17-17	1.04	511	<1	0.07	27	310	8	0.54	<2	6	10	<20	0.13	<10	<10	46	<10	67
		LX17-19	1.38	463	1	0.02	46	400	4	0.31	<2	4	5	<20	0.11	<10	<10	37	<10	110
Average			1.21	487	1	0.05	37	355	6	0.43	<2	5	8	<20	0.12	<10	<10	42	<10	89
HC7	Median-S GW	LX17-16	0.58	372	1	0.04	15	520	2	0.03	<2	3	15	<20	0.09	<10	<10	29	<10	34
HC8	High-S GW	LX17-20	1.43	518	<1	0.02	38	340	3	0.33	<2	4	4	<20	0.14	<10	<10	36	<10	105
		LX17-21	0.89	359	1	0.05	25	260	33	0.45	<2	6	7	<20	0.15	<10	<10	47	<10	66
Average			1.16	439	1	0.04	32	300	18	0.39	<2	5	6	<20	0.15	<10	<10	42	<10	86

**Notes:**

For humidity cells made up of two samples, a 1:1 mixture is assumed

Values in grey italics are below the analytical detection limit.

**Appendix 4-3, Table 3: Solid Phase Elemental Analysis Results - Overburden**

<b>Sample ID</b>	<b>Ag</b>	<b>Al</b>	<b>As</b>	<b>B</b>	<b>Ba</b>	<b>Be</b>	<b>Bi</b>	<b>Ca</b>	<b>Cd</b>	<b>Co</b>	<b>Cr</b>	<b>Cu</b>
	<i>ppm</i>	<i>%</i>	<i>ppm</i>	<i>ppm</i>	<i>ppm</i>	<i>ppm</i>	<i>ppm</i>	<i>%</i>	<i>ppm</i>	<i>ppm</i>	<i>ppm</i>	<i>ppm</i>
LX-BDT-01	0.2	3.17	225	<10	50	0.8	<2	0.06	<0.5	13	37	27
LX-BDT-02	<0.2	1.5	24	<10	30	<0.5	<2	0.04	<0.5	5	21	16
LX-BDT-03	0.2	0.98	54	<10	50	<0.5	<2	0.13	<0.5	5	14	12
LX-BDT-04	<0.2	0.53	15	<10	30	<0.5	<2	0.02	<0.5	2	8	4
LX-BDT-05	0.2	2.39	105	<10	50	0.9	<2	0.05	<0.5	10	30	15
<b>Sample ID</b>	<b>Fe</b>	<b>Ga</b>	<b>Hg</b>	<b>K</b>	<b>La</b>	<b>Mg</b>	<b>Mn</b>	<b>Mo</b>	<b>Na</b>	<b>Ni</b>	<b>P</b>	<b>Pb</b>
	<i>%</i>	<i>ppm</i>	<i>ppm</i>	<i>%</i>	<i>ppm</i>	<i>%</i>	<i>ppm</i>	<i>ppm</i>	<i>%</i>	<i>ppm</i>	<i>ppm</i>	<i>ppm</i>
LX-BDT-01	3.88	10	0.064	0.25	20	0.95	490	1	<0.01	33	440	13
LX-BDT-02	2.63	10	0.061	0.13	10	0.32	245	1	0.01	17	260	9
LX-BDT-03	1.34	10	0.053	0.19	30	0.32	162	1	0.01	22	220	10
LX-BDT-04	0.96	<10	0.038	0.1	<10	0.1	112	<1	0.01	4	110	7
LX-BDT-05	4.51	10	0.054	0.16	20	0.72	441	1	0.01	19	330	9
<b>Sample ID</b>	<b>S</b>	<b>Sb</b>	<b>Sc</b>	<b>Sr</b>	<b>Th</b>	<b>Ti</b>	<b>Tl</b>	<b>U</b>	<b>V</b>	<b>W</b>	<b>Zn</b>	
	<i>%</i>	<i>ppm</i>	<i>ppm</i>	<i>ppm</i>	<i>ppm</i>	<i>%</i>	<i>ppm</i>	<i>ppm</i>	<i>ppm</i>	<i>ppm</i>	<i>ppm</i>	
LX-BDT-01	0.01	<2	4	5	<20	0.09	<10	<10	39	<10	74	
LX-BDT-02	0.02	<2	2	8	<20	0.05	<10	<10	27	<10	28	
LX-BDT-03	0.08	<2	1	9	<20	0.04	<10	<10	13	<10	37	
LX-BDT-04	<0.01	<2	1	4	<20	0.04	<10	<10	16	<10	10	
LX-BDT-05	0.01	<2	3	6	<20	0.08	<10	<10	45	20	64	

**Note:** Values in grey italics are below the analytical detection limit.

**Appendix 4-3, Table 4: Solid Phase Elemental Analysis Results - Haul Road**

Sample ID	Ag	Al	As	B	Ba	Be	Bi	Ca	Cd	Co	Cr	Cu
	ppm	%	ppm	ppm	ppm	ppm	ppm	%	ppm	ppm	ppm	ppm
LX-BDR-01	<0.2	1.03	21	<10	60	<0.5	<2	0.14	<0.5	4	26	9
LX-BDR-02	<0.2	1.69	<2	<10	110	<0.5	<2	0.14	<0.5	13	30	20
LX-BDR-03	<0.2	1.1	2	<10	70	<0.5	<2	0.17	<0.5	7	26	17
LX-BDR-04	<0.2	1.21	<2	<10	120	<0.5	<2	0.12	<0.5	8	29	11
LX-BDR-05	<0.2	2.79	18	<10	190	0.5	<2	0.78	<0.5	13	52	28
<hr/>												
Sample ID	Fe	Ga	Hg	K	La	Mg	Mn	Mo	Na	Ni	P	Pb
	%	ppm	ppm	%	ppm	%	ppm	ppm	%	ppm	ppm	ppm
LX-BDR-01	1.92	<10	<0.005	0.38	10	0.5	212	1	0.02	15	380	3
LX-BDR-02	2.87	10	<0.005	0.92	20	0.95	392	1	0.03	31	610	5
LX-BDR-03	2.08	<10	<0.005	0.56	10	0.65	302	1	0.02	18	550	7
LX-BDR-04	2.12	<10	<0.005	0.77	10	0.6	336	<1	0.03	21	440	<2
LX-BDR-05	3.63	10	<0.005	1.17	30	1.13	625	<1	0.11	38	1130	7
<hr/>												
Sample ID	S	Sb	Sc	Sr	Th	Ti	Tl	U	V	W	Zn	
	%	ppm	ppm	ppm	ppm	%	ppm	ppm	ppm	ppm	ppm	
LX-BDR-01	<0.01	<2	2	12	<20	0.1	<10	<10	21	<10	38	
LX-BDR-02	<0.01	3	3	10	<20	0.13	<10	<10	32	<10	62	
LX-BDR-03	<0.01	<2	2	15	<20	0.11	<10	<10	25	<10	45	
LX-BDR-04	<0.01	<2	3	7	<20	0.14	<10	<10	28	<10	39	
LX-BDR-05	<0.01	<2	7	55	<20	0.18	<10	<10	53	<10	66	

Note: Values in grey italics are below the analytical detection limit.

## ***Appendix 4-4: Particle Size Distribution Results***

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ATLANTIC GOLD



## ***Appendix 4-5: Kinetic Test Results***

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ATLANTIC GOLD

**Appendix 4-5, Table 1: Kinetic Test Results - Cell Description**

Cell No.	Sample ID	Sample Type	Method Reference	Column Dimensions		Column Packing			Total Volume of Initial Flushings	Flushing Rate/Weekly Input*	Temp	Sampling Frequency	Start-up Date	Sampling Day	Operation Procedure	Sample Prep for Flushings
				Inner Diameter (cm)	Length (cm)	Dry Wt. of Sample (kg)	Other Materials Used	Column Material	(mL)	(mL)	(°C)		2018			
HC 1	LX17-12	Waste Rock	MEND	10.00	20.00	1.00	Plexiglas perforated disk & nylon mesh	Plexiglas	500	500	20-22 °C	Weekly	26-Apr	Thursday	Flood Leach	None
HC 2	LX17-24 + LX17-22 Composite	Waste Rock	MEND	10.00	20.00	1.00	Plexiglas perforated disk & nylon mesh	Plexiglas	500	500	20-22 °C	Weekly	26-Apr	Thursday	Flood Leach	None
HC 3	LX17-10 + LX17-02 Composite	Waste Rock	MEND	10.00	20.00	1.00	Plexiglas perforated disk & nylon mesh	Plexiglas	500	500	20-22 °C	Weekly	26-Apr	Thursday	Flood Leach	None
HC 4	LX17-03 + LX17-04 Composite	Waste Rock	MEND	10.00	20.00	1.00	Plexiglas perforated disk & nylon mesh	Plexiglas	500	500	20-22 °C	Weekly	26-Apr	Thursday	Flood Leach	None
HC 5	LX17-23 + LX17-01 Composite	Waste Rock	MEND	10.00	20.00	1.00	Plexiglas perforated disk & nylon mesh	Plexiglas	500	500	20-22 °C	Weekly	26-Apr	Thursday	Flood Leach	None
HC 6	LX17-17 + LX17-19 Composite	Waste Rock	MEND	10.00	20.00	1.00	Plexiglas perforated disk & nylon mesh	Plexiglas	500	500	20-22 °C	Weekly	26-Apr	Thursday	Flood Leach	None
HC 7	LX17-16	Waste Rock	MEND	10.00	20.00	1.00	Plexiglas perforated disk & nylon mesh	Plexiglas	500	500	20-22 °C	Weekly	26-Apr	Thursday	Flood Leach	None
HC 8	LX17-20 + LX17-21 Composite	Waste Rock	MEND	10.00	20.00	1.00	Plexiglas perforated disk & nylon mesh	Plexiglas	500	500	20-22 °C	Weekly	26-Apr	Thursday	Flood Leach	None

































## ***Appendix 4-6: Shake Flask Extraction Results***

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ATLANTIC GOLD

### Appendix 4-6: Shake Flask Extraction Results - Overburden

Parameters	Units	CCME WQG		LX-BDT-01	LX-BDT-02	LX-BDT-03	LX-BDT-04	LX-BDT-05
		Short Term	Long Term	L2178440-1	L2178440-2	L2178440-3	L2178440-4	L2178440-5
<i>Physical Tests (Soil)</i>								
Moisture	%	-	-	<0.25	<0.25	<0.25	<0.25	<0.25
<i>Leachable Anions &amp; Nutrients (Soil)</i>								
Alkalinity, Total (as CaCO <sub>3</sub> )	mg/L	-	-	<2.0	<2.0	<2.0	<2.0	<2.0
Conductivity	µS/cm	-	-	12.6	15.7	89.1	36.7	19.5
pH	pH	6.5-9	-	5.27	5.15	4.88	4.54	5.83
Sulfate (SO <sub>4</sub> )	mg/L	-	-	2.24	1.23	10.1	1.42	2.08
<i>Leachable Metals (Soil)</i>								
Aluminum (Al)-Leachable	mg/L	-	0.1	<b>0.108</b>	<b>0.683</b>	<b>1.51</b>	<b>1.26</b>	<b>0.483</b>
Antimony (Sb)-Leachable	mg/L	-	-	<0.00010	<0.00010	0.00015	<0.00010	<0.00010
Arsenic (As)-Leachable	mg/L	-	0.005	0.0012	0.0012	<b>0.315</b>	0.0038	<b>0.0084</b>
Barium (Ba)-Leachable	mg/L	-	-	0.0044	0.0040	0.0271	0.0066	0.0097
Beryllium (Be)-Leachable	mg/L	-	-	<0.00050	<0.00050	<0.00050	<0.00050	<0.00050
Bismuth (Bi)-Leachable	mg/L	-	-	<0.00050	<0.00050	<0.00050	<0.00050	<0.00050
Boron (B)-Leachable	mg/L	29	1.5	<0.010	<0.010	0.015	<0.010	<0.010
Cadmium (Cd)-Leachable	mg/L	0.001	0.00009	<0.000050	<0.000050	<b>0.000217</b>	0.000075	<0.000050
Calcium (Ca)-Leachable	mg/L	-	-	0.87	0.59	8.78	1.03	1.18
Chromium (Cr)-Leachable	mg/L	-	0.001	<0.00050	0.00088	<b>0.00118</b>	<b>0.00126</b>	0.00064
Cobalt (Co)-Leachable	mg/L	-	-	0.00084	0.00042	0.00761	0.00054	0.00101
Copper (Cu)-Leachable	mg/L	-	0.002	<0.0010	<0.0010	<b>0.0085</b>	0.0014	<b>0.0033</b>
Iron (Fe)-Leachable	mg/L	-	0.3	<0.030	0.194	<b>0.425</b>	<b>0.490</b>	0.262
Lead (Pb)-Leachable	mg/L	-	0.001	<0.00010	0.00013	0.00090	<b>0.00129</b>	0.00012
Lithium (Li)-Leachable	mg/L	-	-	<0.0050	<0.0050	<0.0050	<0.0050	<0.0050
Magnesium (Mg)-Leachable	mg/L	-	-	0.237	0.345	1.44	0.556	0.471
Manganese (Mn)-Leachable	mg/L	-	-	0.107	0.0446	0.131	0.0681	0.106
Mercury (Hg)-Leachable	mg/L	-	0.000026	<0.000050	<0.000050	<0.000050	<0.000050	<0.000050
Molybdenum (Mo)-Leachable	mg/L	-	0.073	<0.00010	<0.00010	0.00023	<0.00010	<0.00010
Nickel (Ni)-Leachable	mg/L	-	0.025	0.00098	0.00118	<b>0.0360</b>	0.00151	0.00179
Phosphorus (P)-Leachable	mg/L	-	-	<0.30	<0.30	<0.30	<0.30	<0.30
Potassium (K)-Leachable	mg/L	-	-	0.286	0.551	2.92	1.72	0.793
Selenium (Se)-Leachable	mg/L	-	0.001	<0.00050	0.00068	0.00079	<b>0.00101</b>	0.00085
Silicon (Si)-Leachable	mg/L	-	-	3.44	1.37	1.52	0.752	1.71
Silver (Ag)-Leachable	mg/L	-	0.00025	<0.000050	<0.000050	<0.000050	<0.000050	<0.000050
Sodium (Na)-Leachable	mg/L	-	-	0.258	0.629	2.65	1.24	0.757
Strontium (Sr)-Leachable	mg/L	-	-	0.00454	0.00563	0.0403	0.00885	0.0105
Thallium (Tl)-Leachable	mg/L	-	0.0008	<0.00010	<0.00010	<0.00010	<0.00010	<0.00010
Tin (Sn)-Leachable	mg/L	-	-	<0.00050	<0.00050	<0.00050	<0.00050	<0.00050
Titanium (Ti)-Leachable	mg/L	-	-	<0.010	<0.010	0.016	0.027	0.011
Uranium (U)-Leachable	mg/L	0.033	0.015	0.000020	0.000087	0.000101	0.000073	0.000103
Vanadium (V)-Leachable	mg/L	-	-	<0.0010	<0.0010	0.0038	0.0023	<0.0010
Zinc (Zn)-Leachable	mg/L	0.037	0.007	<0.010	<0.010	<b>0.045</b>	<b>0.012</b>	<0.010

**Notes:**

Aluminum guideline is based on pH > 6.5

Hardness dependent guidelines (Cu, Pb, Ni) are based on a hardness of 10 mg/L

Grey shading indicates values above the short-term CCME water quality guideline

Bold italic text indicates values above the long-term CCME water quality guideline